

Chemistry Unit 8 Worksheet 4

Examples of Every Kind of Problem

This is a separate sheet of paper, so this is a complete solution to each of the problems. You can follow the problems without an answer. We want to make your final solution.

1. Calculate the number of moles of potassium chlorate, KClO_3 , that must decompose to produce potassium chloride, KCl , and 1.5 moles of oxygen gas.



Substance	KClO_3	KCl	O_2
Moles	1.5	0	1.5
Change	-1.5	+1.5	+1.5
Initial	0	0	0

1.5 moles KClO_3 must decompose

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 0.1 g of magnesium?



Substance	Mg	HCl	MgCl_2	H_2
Moles	0.004	0	0	0
Change	-0.004	-0.008	+0.004	+0.004
Initial	0	0	0	0

0.008 moles HCl needed

3. Ethane, C_2H_6 , reacts with oxygen gas to produce carbon dioxide gas and water vapor. What number of moles of oxygen gas is required to react with 0.1 moles of ethane?



Substance	C_2H_6	O_2	CO_2	H_2O
Moles	0.1	0	0	0
Change	-0.1	-0.35	+0.2	+0.3
Initial	0	0	0	0

0.35 moles O_2 must react with 0.1 moles C_2H_6 to react

4. Sodium metal reacts with sulfuric acid to produce sodium sulfate and hydrogen gas. How many moles of sulfuric acid are needed to react with 0.1 moles of sodium metal?



Substance	Na	H_2SO_4	Na_2SO_4	H_2
Moles	0.1	0	0	0
Change	-0.1	-0.05	+0.05	+0.05
Initial	0	0	0	0

0.05 moles H_2SO_4 must react with 0.1 moles Na to react

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Victor M. Corman



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