



c) [NaOH] (in mol/L),

$$x \text{ mol NaOH} = 0.02500 \text{ L} \times \frac{0.00690 \text{ mol}}{\text{L}} \times \frac{1 \text{ mol NaOH}}{1 \text{ mol HA}} = 1.73 \times 10^{-4} \text{ mol NaOH}$$

$$[NaOH] = \frac{1.73 \times 10^{-4} \text{ mol}}{0.02135 \text{ L}} = \underline{\underline{0.0081 \text{ M}}}$$

d) if the resulting solution is acidic, basic or neutral at the equivalence point, and

A^- is weakly basic \Rightarrow basic solution

e) an appropriate indicator for the titration.

Phenolphthalein

8. Triethanolamine, a weak monoprotic base, is titrated with HCl. The following data are obtained.

- 20.00 mL of triethanolamine
- 16.30 mL of HCl
- Initial pH of acid = 10.355

– $pH_{\frac{1}{2}} = 7.763$ at addition of 8.15 mL of acid during titration

$$pOH_{\frac{1}{2}} = 14 - 7.763 = 6.237$$

Determine

a) the K_b for triethanolamine,

$$K_b = [OH^-]^2$$

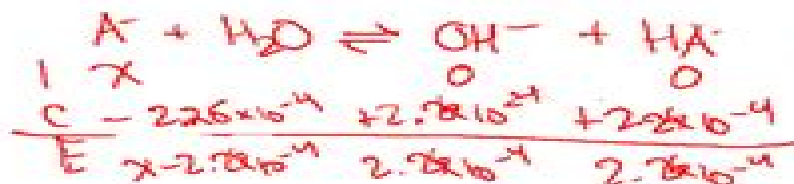
$$= 10^{-pOH_{\frac{1}{2}}}$$

$$= \underline{\underline{5.79 \times 10^{-7}}}$$

THIS should read initial [triethanolamine] \star

b) initial ~~10.355~~ (in mol/L).

$$pOH_{\text{init}} = 14.000 - 10.355 = 3.645$$



$$K_a = \frac{[OH^-][HA]}{[CH_3COOH]} = 5.79 \times 10^{-7}$$

$$x - 2.2 \times 10^{-4} = \frac{(2.26 \times 10^{-4})^2}{5.79 \times 10^{-7}}$$

$$x = 0.0859 + 0.000212$$

$$\underline{\underline{x = 0.0861 \text{ M}}}$$

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