

## Chemistry Final Exam Review Questions

### Useful Values and Equations

Some of the following values and equations may be useful for answering some of the questions and problems.

$$Q_F = 1.84 Q_C + 12$$

speed of light =  $3.00 \times 10^8$  m/s

2000

$R = 0.082057 \text{ L-atm / deg-mole}$

Planck's constant =  $6.63 \times 10^{-34} \text{ J.s}$

15. *Journal of the American Medical Association*, 277:1033-1034 (1997).

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- One of the following does **not** describe gases. Which one is it?
  - Expands infinitely.
  - Rigid shape.
  - Easily compressed.
  - Relatively low densities.
- The following properties describe the element, mercury (Hg). Which one is a **chemical** property?
  - Its density is  $5.43 \text{ g/cm}^3$ .
  - It is a silvery-white metal.
  - It is a liquid at 298 K.
  - Mercury(II) sulfide forms when it is exposed to sulfide ions.
- All of the following are properties of  $\text{SO}_2$ . Which one is a **physical** property?
  - It reacts with oxygen to form  $\text{SO}_3^{2-}$ .
  - It condenses to a colorless liquid at  $14^\circ\text{F}$  and atmospheric pressure.
  - It is a strong reducing agent.
  - It reacts with water and oxygen to form an acid.
- Which of the following numbers has 4 significant figures?
  - 0.04309
  - 0.0430
  - 0.0431
  - 0.43980
- How many significant figures should there be in the answer to the following math equation?
$$2.834 + 5.71520 + 2.12 + 178.1 + 250.2619$$
  - 3
  - 4
  - 5
  - 6
- How many significant figures should there be in the answer to the following math equation?
$$23.413 \text{ g} + (2.15 \text{ cm} \times 1.1 \text{ cm} \times 3.73 \text{ cm})$$
  - 2
  - 3
  - 4
  - 5
- At  $25^\circ\text{C}$ , one  $\text{cm}^3$  of iron has a mass of 7.62 grams. What volume of iron is occupied by 125 grams of iron?
  - $0.82 \text{ cm}^3$
  - $8.2 \text{ cm}^3$
  - $0.164 \text{ cm}^3$
  - $16.4 \text{ cm}^3$
- A cube of lead has a volume of  $6.61 \text{ cm}^3$ . Its mass is 75.0 grams. What is its density?
  - $3.85 \text{ g/cm}^3$
  - $7.41 \text{ g/cm}^3$
  - $11.3 \text{ g/cm}^3$
  - $22.6 \text{ g/cm}^3$

# Modeling Chemistry Review For Final Exam

**Wenbin Ji**



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*Matriculation model answers: Latin. Being the papers from June 1888 to Jan. 1891 (to Jan. 1894, Jan. 1892 to June 1897, June 1894 to Jan. 1900, Jan. 1896 to June 1901, to Sept. 1902, to Jan. 1904, Sept. 1902 to Jan. 1906, to Jan. 1907, to Jan. 1908, to Jan. 1909, to Jan. 1910, to June 1911, Jan. 1907 to Jan. 1912, [June 1912 to June 1915]), with answers by B.J. Hayes and W.F. Masom (with answers by tutors of University correspondence college).* London univ, exam. papers,1891

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