

## Chemistry Final Exam Review Questions

### Useful Values and Equations

Some of the following values and equations may be useful for answering some of the questions and problems.

$$Q_F = 1.84 Q_C + 12$$

speed of light =  $3.00 \times 10^8$  m/s

2000

$R = 0.082057 \text{ L-atm / deg-mole}$

Planck's constant =  $6.63 \times 10^{-34} \text{ J.s}$

15. *Journal of the American Medical Association*, 277:1005-1006 (1997).

### Multiple Choice

Identify the letter of the choice that best completes the statement or answers the question.

- One of the following does **not** describe gases. Which one is it?
  - Expands infinitely.
  - Rigid shape.
  - Easily compressed.
  - Relatively low densities.
- The following properties describe the element, mercury (Hg). Which one is a **chemical** property?
  - Its density is  $5.43 \text{ g/cm}^3$ .
  - It is a silvery-white metal.
  - It is a liquid at 298 K.
  - Mercury(II) sulfide forms when it is exposed to sulfide ions.
- All of the following are properties of  $\text{SO}_2$ . Which one is a **physical** property?
  - It reacts with oxygen to form  $\text{SO}_3^{2-}$ .
  - It condenses to a colorless liquid at  $14^\circ\text{F}$  and atmospheric pressure.
  - It is a strong reducing agent.
  - It reacts with water and oxygen to form an acid.
- Which of the following numbers has 4 significant figures?
  - 0.04309
  - 0.0430
  - 0.0431
  - 0.43980
- How many significant figures should there be in the answer to the following math equation?
$$2.834 + 5.71520 + 2.12 + 178.1 + 250.2619$$
  - 3
  - 4
  - 5
  - 6
- How many significant figures should there be in the answer to the following math equation?
$$23.413 \text{ g} + (2.15 \text{ cm} \times 1.1 \text{ cm} \times 3.73 \text{ cm})$$
  - 2
  - 3
  - 4
  - 5
- At  $25^\circ\text{C}$ , one  $\text{cm}^3$  of iron has a mass of 7.62 grams. What volume of iron is occupied by 125 grams of iron?
  - $0.82 \text{ cm}^3$
  - $8.2 \text{ cm}^3$
  - $0.164 \text{ cm}^3$
  - $16.4 \text{ cm}^3$
- A cube of lead has a volume of  $6.61 \text{ cm}^3$ . Its mass is 75.0 grams. What is its density?
  - $3.85 \text{ g/cm}^3$
  - $7.41 \text{ g/cm}^3$
  - $11.3 \text{ g/cm}^3$
  - $22.6 \text{ g/cm}^3$

# Modeling Chemistry Final Exam Review Answers

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