

## Chemistry Unit 8 Worksheet 4

### Examples of Every Kind of Problem

This is a separate sheet of paper, so this is a complete solution to each of the problems. You can follow the problems without an answer. We want to make your final solution.

12. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$ , that must decompose to produce potassium chloride  $\text{KCl}$  and a total of 1.5 moles of oxygen gas.



Substance	$\text{KClO}_3$	$\text{KCl}$	$\text{O}_2$
Moles	1.0	0	1.5
Change	-1.0	+2.0	+1.5
Initial	0	0	0

**1.0 moles  $\text{KClO}_3$  must**

13. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 5.0 g of magnesium?



Substance	$\text{Mg}$	$\text{HCl}$	$\text{MgCl}_2$	$\text{H}_2$
Moles	0.190	0	0	0
Change	-0.190	+0.380	+0.190	+0.190
Initial	0	0	0	0

**0.380 moles HCl needed**

14. Ethane,  $\text{C}_2\text{H}_6$ , reacts with oxygen gas to produce carbon dioxide gas and water vapor. What number of moles of oxygen gas is required to react with 0.50 moles of ethane?



Substance	$\text{C}_2\text{H}_6$	$\text{O}_2$	$\text{CO}_2$	$\text{H}_2\text{O}$
Moles	0.50	0	0	0
Change	-0.50	+0.875	+1.00	+1.50
Initial	0	0	0	0

**0.875 moles  $\text{O}_2$  must be used to react with 0.50 moles  $\text{C}_2\text{H}_6$**

15. Sodium metal reacts with sulfur dioxide to produce sodium sulfide and sodium sulfate. How many moles of sodium metal are needed to produce 1.00 mole of sodium sulfate?



Substance	$\text{Na}$	$\text{SO}_2$	$\text{Na}_2\text{S}$	$\text{Na}_2\text{SO}_4$
Moles	0.00	0.00	0	0
Change	-4.00	-1.00	+1.00	+2.00
Initial	0	0	0	0

**4.00 moles Na must be used to produce 2.00 moles  $\text{Na}_2\text{SO}_4$**

# Modeling Chemistry Unit 11 Ws1

**Arthur James Wells**



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