Equilibrium HOMEWORK

Lesson 1:

1. Write the equilibrium constant expressions for the following reactions. Use pressures for gases and concentrations for solutions.

   a) $NaOH(s) \leftrightarrow Na^+(aq) + OH^-(aq)$ $K_c = [Na^+][OH^-]$

   b) $4CH_3OH(g) + 7O_2(g) \leftrightarrow 4H_2O(g) + 2CO_2(g)$ $K_c = \frac{[H_2O]^4[CO_2]^2}{[CH_3OH]^4[O_2]^7}$ $K_p = \frac{P_{H_2O}^4 P_{CO_2}^2}{P_{CH_3OH}^4 P_{O_2}^7}$

   c) $2CO_2(g) \leftrightarrow CO_4(s)$ $K_c = \frac{1}{[CO_2]^2}$

   *Remember that SOLIDS and LIQUIDS SHOULD not be included in the equilibrium expression!

   d) $H_2O(l) \leftrightarrow H_2O(g)$ $K_c = [H_2O]$ $K_p = P_{H_2O}$

   e) $PbCO_3(aq) + 2HNO_3(l) \leftrightarrow 2NaO^+(aq) + CO_4^{2-}(aq)$ $K_c = \frac{[NaO^+]^2[CO_4^{2-}]}{[PbCO_3][HNO_3]^2}$

   f) $(NH_4)_2S(aq) \leftrightarrow 2NH_4^+(aq) + S^{2-}(aq)$ $K_c = [NH_4^+]^2[S^{2-}]$

2. The reaction below takes place in a 1-Liter container. It produces a mixture that contains 0.450M $(NH_4)_2S$, 0.220M NH_4^+, and 0.108M S^{2-}. This mixture is at equilibrium at 1200 K.

   $(NH_4)_2S(aq) \leftrightarrow 2NH_4^+(aq) + S^{2-}(aq)$

   a. What is the equilibrium constant for this reaction at 1200 K?

   $K_c = \frac{[NH_4^+]^2[S^{2-}]}{[(NH_4)_2S]}$

   $K_c = \frac{(0.220)^2(0.108)}{(0.450)} = 0.0116$

   b. Based on the equilibrium constant, is this reaction reactant-favored or product-favored at 1200 K?

   reactant-favored, because K_c is less than one, meaning there are more reactants than products at equilibrium.

3. The reaction below takes place in a 1-Liter container. It produces a gaseous mixture that contains 0.30atm CO, 0.60 atm H_2, 0.20mol H_2O, and 0.075 atm CH_4. This mixture is at equilibrium at 1200 K.

   $CH_4(g) + H_2O(l) \leftrightarrow CO(g) + 3H_2(g)$

   a. What is the equilibrium constant for this reaction at 1200 K?

   *Don't include liquid water in the equilibrium constant!

   $K_c = \frac{[CO][H_2]^3}{[CH_4]}$ $K_p = \frac{P_{CO}P_{H_2}^3}{P_{CH_4}}$

   $K_c = \frac{(0.30)(0.60)^3}{(0.075)} = 0.432$ $K_p = 0.432$

   b. Based on the equilibrium constant, is this reaction reactant-favored or product-favored at 1200 K?

   reactant-favored, because K_c is less than one, meaning there are more reactants than products at equilibrium.

4. The reaction below takes place in a 50.0-Liter container. It produces a gaseous mixture that contains 0.200mol N_2, 0.600mol H_2, and 1.000mol NH_3. This mixture is at equilibrium at 1200 K.

   $N_2(g) + 3H_2(g) \leftrightarrow 2NH_3(g)$

   a. What is the equilibrium constant for this reaction at 1200 K?

   *Because the stoichiometric ratios in the SOLUTE and COEFFICIENTS are equivalent as SIMPLIFIED RATIOS, the K_c and K_p are equivalent!

   $K_c = \frac{[NH_3]^2}{[N_2][H_2]^3}$ $K_p = \frac{P_{NH_3}^2}{P_{N_2}P_{H_2}^3}$

   *Remember that GASEOUS reactants are measured using an equilibrium expression. If given the GASEOUS amounts in moles, divide by Liters!

   $K_c = \frac{(0.020)^2}{(0.004)(0.012)^3} = 5.787 \times 10^4$ $K_p = 5.79 \times 10^4$

   b. Based on the equilibrium constant, is this reaction reactant-favored or product-favored at 1200 K?

   product-favored, because K_c or K_p is greater than one, meaning there are more products than reactants at equilibrium.

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Frank M. Dunnivant,Elliot Anders

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