

MIXED GAS LAWS WORKSHEET

Key

Directions: Examine each question and then write the form of the gas law you plan to use to solve each question. Show which values you have, which values are missing and/or which values need to be calculated. Be careful to use standard units of volume (liters), temperature (Kelvins) & pressure (Atm or mm of Hg).

Note: 1 atm = 760 mm Hg

1. A gas occupies 3.5L at 2.5 mm Hg pressure. What is the volume at 10 mm Hg at the same temperature?

$$P_1 = 2.5 \text{ mmHg} \quad P_2 = 10 \text{ mmHg} \quad P_1 V_1 = P_2 V_2 \quad \frac{(2.5 \text{ mmHg})(3.5 \text{ L})}{10 \text{ mmHg}} = \frac{(10 \text{ mmHg}) V_2}{10 \text{ mmHg}}$$

$$V_1 = 3.5 \text{ L} \quad V_2 = ?$$

$$0.875 \text{ L} = V_2$$

2. A constant volume of oxygen is heated from 100°C to 185°C. The initial pressure is 4.1 atm. What is the final pressure?

$$P_1 = 4.1 \text{ atm} \quad \frac{P_1}{T_1} = \frac{P_2}{T_2} \quad \frac{4.1 \text{ atm}}{373 \text{ K}} = \frac{P_2}{458 \text{ K}}$$

$$T_1 = 100^\circ\text{C} + 273 = 373 \text{ K}$$

$$P_2 = ?$$

$$T_2 = 185^\circ\text{C} + 273 = 458 \text{ K}$$

$$P_2 = \frac{(4.1 \text{ atm})(458 \text{ K})}{373 \text{ K}} = 5.0 \text{ atm}$$

3. A sample of 25L of NH_3 gas at 10°C is heated at constant pressure until it fills a volume of 50L. What is the new temperature in °C?

$$V_1 = 25 \text{ L} \quad \frac{V_1}{T_1} = \frac{V_2}{T_2} \quad T_2 = \frac{(50 \text{ L})(283)}{25 \text{ L}} = 566 \text{ K}$$

$$T_1 = 10^\circ\text{C} + 273 = 283 \text{ K}$$

$$V_2 = 50 \text{ L}$$

$$T_2 = ?$$

$$^\circ\text{C} = \text{K} - 273 = 566 - 273 = 293^\circ\text{C}$$

4. An unknown gas weighs 34g and occupies 6.7L at 2 atm and 245K. What is its molecular weight?

$$\text{Molar Mass} = \frac{g}{\text{mol}} \quad PV = nRT$$

$$(2 \text{ atm})(6.7 \text{ L}) = n (0.0821 \frac{\text{L atm}}{\text{mol K}})(245 \text{ K})$$

$$n = \frac{(2 \text{ atm})(6.7 \text{ L})}{(0.0821 \frac{\text{L atm}}{\text{mol K}})(245 \text{ K})} = 0.666 \text{ mol}$$

$$\text{Molar mass} = \frac{34 \text{ g}}{0.666 \text{ mol}} = 51 \text{ g/mol}$$

5. An ideal gas occupies 400ml at 270 mm Hg and 65°C. If the pressure is changed to 1.4 atm and the temperature is increased to 100°C, what is the new volume?

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2} \quad \frac{(0.355 \text{ atm})(400 \text{ mL})}{318 \text{ K}} = \frac{(1.4 \text{ atm})(V_2)}{373 \text{ K}}$$

$$V_2 = \frac{(0.355 \text{ atm})(400 \text{ mL})(373 \text{ K})}{(318 \text{ K})(1.4 \text{ atm})} = 112 \text{ mL}$$

$$\text{Unit: } \begin{cases} P_1 = 270 \text{ mmHg} \\ V_1 = 400 \text{ mL} \\ T_1 = 65^\circ\text{C} + 273 = 338 \text{ K} \\ P_2 = 1.4 \text{ atm} \\ V_2 = ? \\ T_2 = 100^\circ\text{C} + 273 = 373 \text{ K} \end{cases}$$

6. What is the volume of 23g of neon gas at 1°C and a pressure of 2 atm?

$$PV = nRT \quad 23 \text{ g Ne} \left(\frac{1 \text{ mol Ne}}{20.18 \text{ g Ne}} \right) = 1.14 \text{ mol Ne}$$

$$V = \frac{nRT}{P} = \frac{(1.14 \text{ mol})(0.0821 \frac{\text{L atm}}{\text{mol K}})(274 \text{ K})}{2 \text{ atm}} = 12.8 \text{ L}$$

Mixed Gas Law Calculations Answers

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