

4.7 CHEMISTRY (12%)

4.7.1 Chemistry Paper 1 (12%)

1. (a) Ether (1 mark)
- (b) Add water to each compound.
Add Na_2CO_3 or NaHCO_3 to a suitable part of each sample.
 $\text{C}_2\text{H}_5\text{CO}$ no effervescence (1 mark)
 $\text{C}_2\text{H}_5\text{CO}_2$ effervescence (1 mark)
 no odd smell (propionate ethyl ester) and water
 $\text{C}_2\text{H}_5\text{CO}$ water layer starts to grow
 $\text{C}_2\text{H}_5\text{CO}_2$ no change
 Or add acidified Potassium Manganate (VII)
 $\text{C}_2\text{H}_5\text{CO}$ will be decolourised
 $\text{C}_2\text{H}_5\text{CO}_2$ no change
2. (a) Ether (Na_2CO_3) (1 mark)
- (b) — Sodium is very reactive (can electrolyse) (1 mark)
 — More reactive than carbon
3. (a) Ether
 Sodium lamp, used in nuclear reactions
 Sodium vapour, sodium amalgams
 (NaOH), Electrodes and cathodes, etc. (1 mark)
4. (a) Endothermic change, when one mole of crystal lattice is broken into its ions in gaseous state. (1 mark)
- (b) Endothermic reaction (percent) (1 mark)
5. (a) Boyle's law: The volume of a fixed mass of a gas is inversely proportional to its pressure or the product of pressure and volume is constant at a fixed temperature. (1 mark)

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l when the half reactions listed above are combined into a balanced chemical equation the ratio of the coefficients $\text{Cu} + \text{MnO}_4^{-}$
4 a a q is

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stoichiometric coefficient of HSO_4^{-} if the above equation is balanced in an acidic medium stuck use a hint report a problem
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redox reaction in exercise 9 by the half reaction method by assuming a basic solution

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