

19 This is a question about catalysis.

A spectacular demonstration of catalytic oxidation is the addition of chromium(III) oxide to ammonia gas in the presence of oxygen. This produces flashes, sometimes described as 'fireflies'.

Some concentrated ammonia solution is allowed to vaporise in a very large flask. Heated chromium(III) oxide catalyst is added.

The equation below shows one possible reaction.



- (a) The nitrogen atoms in ammonia are oxidized. Give the oxidation numbers of the nitrogen atoms in ammonia and nitrogen monoxide.

(2)

NH₃ _____ NO _____

- (b) Nitrogen monoxide is an example of a free radical.

- (i) Explain why nitrogen monoxide is classed as a free radical.

(1)

- (ii) Draw the dot and cross diagram of nitrogen monoxide, using dots (•) for the nitrogen electrons and crosses (x) for the oxygen electrons. Show outer shell electrons only.

(2)

- (c) Suggest **one** suitable safety precaution for this demonstration. Justify your choice. You can assume that the demonstrator is wearing a lab coat and safety goggles.

(2)



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