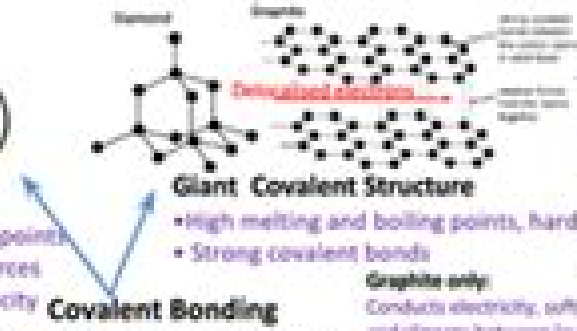
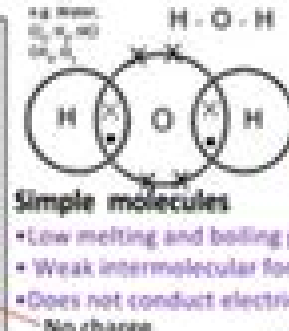




Periodic Table of the Elements

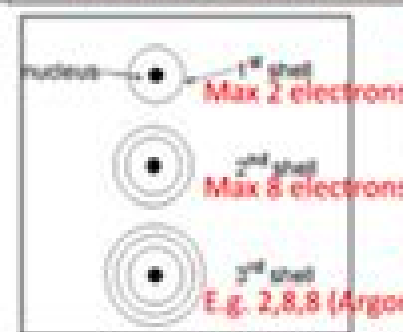
- metals
- non-metals
- transition metals
- alkali metals
- alkali earth metals
- transition metals
- lanthanoids
- actinoids
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- noble gases
- metalloids

An atom:
 Number of Protons = Number of Electrons
 Mass number = Protons + Neutrons

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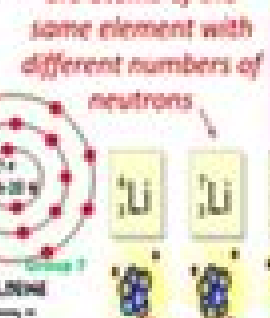
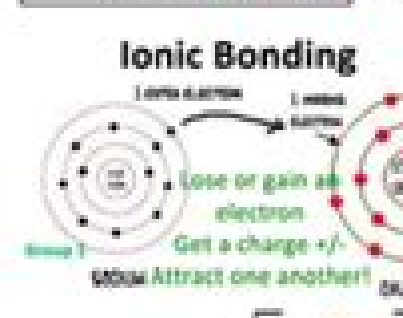
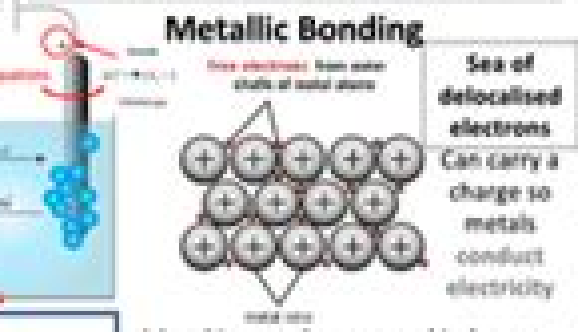
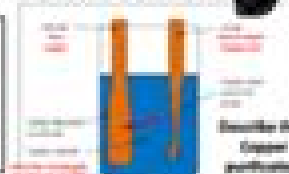
Periodic table arranges elements by **ATOMIC NUMBER** (proton number/the small one)

The number of outer shell electrons match the group the element is found in.
 E.g. Lithium 2,1 is a group 1 element.



ISOTOPES
 are atoms of the same element with different numbers of neutrons

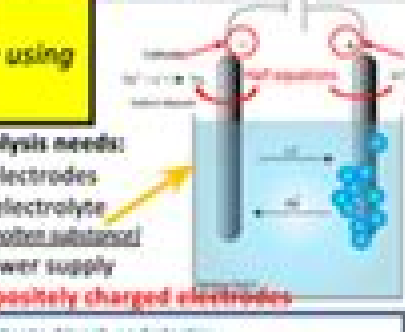
OIL RIG
 Oxidation is loss (of electrons)
 Reduction is gain (of electrons)



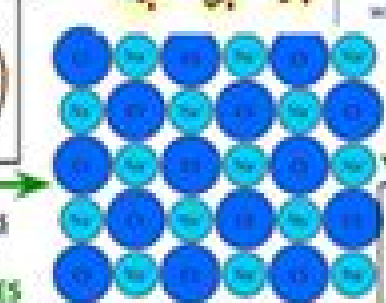
When there are 2 possible products
 The less reactive forms at the electrode

Electrolysis needs:
 • 2 electrodes
 • An electrolyte (ionic or molten substance)
 • A power supply

Ions move toward oppositely charged electrodes



Always metals with non-metals
 + ION and - ION attracted by **STRONG ELECTROSTATIC FORCES**



As a solid: High melting and Boiling Point
 Conducts electricity when melted or dissolved:
 Ions move freely and carry a charge

NANO-TECHNOLOGY
 Means REALLY REALLY REALLY small!!!!

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