

Name Tara ShettyDate 9/24/15 Pd 5

Chemistry – Unit 4 Worksheet 3

Use the following information about the masses of elements in each pair of compounds to help you suggest formulas that account for these ratios.

1. Compounds of carbon and oxygen

Compound A: 57.1 g O / 42.9 g C

Compound B: 72.7 g O and 27.3 C

$$A = \frac{57.1}{42.9} = 1 \frac{1}{3}$$

$$B = \frac{72.7}{27.3} = 2 \frac{2}{3}$$

a. Determine the value of the ratio $\frac{\text{mass O}}{\text{mass C}}$ in each compound. A $1 \frac{1}{3}$ B $2 \frac{2}{3}$

b. How does the mass ratio for compound B compare to that in compound A?

compound B is larger than compound A

c. Express these ratios as improper fractions. $\frac{57.1}{42.9} + \frac{72.7}{27.3} \left(\frac{4}{3} + \frac{8}{3} \right)$

d. For each hypothesis, sketch particle diagrams for the compounds of A and B that account for these mass ratios. Write the formula for the compound in each diagram.

Hypothesis 1

Atoms of C and O have the same mass

A $\frac{4}{3}$ B $\frac{8}{3}$ 

Hypothesis 2

Atoms of O are heavier than C atoms by the ratio in compound A.

A



B



Modeling Chemistry Unit 10 2 Answer

Jessica J Manson



Modeling Chemistry Unit 10 2 Answer:

Selected Solutions for Chemistry, Concepts and Models by Robinson, Odom, and Holtzclaw John H. Meiser, 1992

Chemical Modelling Alan Hinchliffe, 2008-11-19 Chemical Modelling Applications and Theory comprises critical literature reviews of molecular modelling both theoretical and applied Molecular modelling in this context refers to modelling the structure properties and reactions of atoms molecules materials Each chapter is compiled by experts in their fields and provides a selective review of recent literature With chemical modelling covering such a wide range of subjects this Specialist Periodical Report serves as the first port of call to any chemist biochemist materials scientist or molecular physicist needing to acquaint themselves of major developments in the area Volume 5 covers literature published from June 2005 to May 2007

Environmental Organic Chemistry Rene P. Schwarzenbach, Philip M. Gschwend, Dieter M. Imboden, 2016-11-07 Examines in a pedagogical way all pertinent molecular and macroscopic processes that govern the distribution and fate of organic chemicals in the environment and provides simple modeling tools to quantitatively describe these processes and their interplay in a given environmental system Treats fundamental aspects of chemistry physics and mathematical modeling as applied to environmentally relevant problems and gives a state of the art account of the field Teaches the reader how to relate the structure of a given chemical to its physical chemical properties and intrinsic reactivities Provides a holistic and teachable treatment of phase partitioning and transformation processes as well as a more focused and tailor made presentation of physical mathematical and modeling aspects that apply to environmental situations of concern Includes a large number of questions and problems allowing teachers to explore the depth of understanding of their students or allowing individuals who use the book for self study to check their progress Provides a companion website which includes solutions for all problems as well as a large compilation of physical constants and compound properties

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