

Honors Chemistry - Unit 8 Worksheet 1

Mole relationships

For each of the problems below:

- Write the balanced chemical equation.
- Identify what is given (with units) and what you want to find (with units).
- Use coefficients from balanced equation to determine mole ratio.

1. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce sulfur dioxide and water. How many moles of oxygen gas would be needed to completely burn 2.0 moles of hydrogen sulfide?

Equation: $\text{H}_2\text{S}_{(g)}$ + $\text{O}_{2(g)}$ → $\text{SO}_{2(g)}$ + $\text{H}_2\text{O}_{(g)}$

Before:

Change _____

After

2. Aluminum reacts with hydrochloric acid to produce aluminum chloride and hydrogen gas. Write a balanced equation for the reaction. Calculate the number of moles of hydrochloric acid required to react with 0.50 moles of Al.

Equation:

Before:

Change _____

After

3. Ammonia, NH_3 , for fertilizer is made by causing hydrogen and nitrogen to react at high temperature and pressure. How many moles of ammonia can be made from 1.25 moles of nitrogen gas?

Equation:

Before:

Change _____

After

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