

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

Let us suppose that our "paper" writer is a chemist who has to write out the products of the reaction. Show us the products and how to find it. The rest is a matter of neat typed answers.

1. Chlorine gas reacts with a metal to produce metal chloride. MgCl_2 and Mg react. How much chlorine gas is produced per mole of metal? MgCl_2 is 97.9 and 24.31 metal weight.

Equation	$\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$
Moles	1.000
Weight	24.31
Weight	97.9

1.000 mole
97.9 grams

2. How much a displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How much weight of hydrochloric acid is required to produce 10.00 g of magnesium?

Equation	$\text{Mg} + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$			
Moles	1.000	2.000	1.000	1.000
Weight	24.31	72.92	97.9	2.02
Weight	24.31	72.92	97.9	2.02

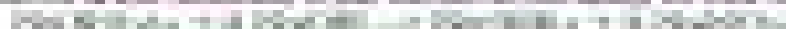
2.000 mole
72.92 grams

3. Chlorine, Cl_2 , reacts with hydrogen gas to produce hydrogen chloride gas. How much gas is required to produce 10.00 g of hydrogen gas by displacement of metal with 10.00 g of metal?

Hydrogen	1.000	1.000	2.000	2.016
Chlorine	1.000	1.000	2.000	70.906
Total	2.000	2.000	4.000	72.922

1.000 mole of Cl_2 is 70.90 g. 1.000 mole of H_2 is 2.02 g.

4. Chlorine gas reacts with hydrogen gas to produce hydrogen chloride gas. How much gas is required to produce 10.00 g of hydrogen gas by displacement of metal with 10.00 g of metal?



2.02 g of H_2 is 1.000 mole

Equation	$2\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$		
Moles	2.000	1.000	2.000
Weight	4.04	70.90	72.92
Weight	4.04	70.90	72.92

1.000 mole of Cl_2 is 70.90 g. 2.000 mole of H_2 is 4.04 g.

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L Darling-Hammond



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