

Chemistry Unit 8 Review

1. The gas silicate, SiOCl_2 , reacts with oxygen gas to produce a gas that is an acidic silicate anhydride.
 If 5.150 grams of SiOCl_2 were consumed using this chemical reaction, what mass of the gas would be produced?



$$5.150 \text{ g SiOCl}_2 \times \frac{1 \text{ mol}}{170.1 \text{ g}} \times \text{stoichiometric factor} \left(\frac{1 \text{ mol Cl}_2}{1 \text{ mol SiOCl}_2} \right) = 0.03025 \text{ mol Cl}_2$$

2. Magnesium sulfate, MgSO_4 , decomposes when heated to produce magnesium sulfate and oxygen gas.
 If 4.12 grams of MgSO_4 were heated in a test tube, what mass of oxygen gas would be produced?



$$4.12 \text{ g MgSO}_4 \times \frac{1 \text{ mol}}{120.3 \text{ g}} \times \text{stoichiometric factor} \left(\frac{1 \text{ mol SO}_3}{1 \text{ mol MgSO}_4} \right) = 0.0342 \text{ mol SO}_3 = 0.0342 \text{ mol SO}_3 \times \frac{80.06 \text{ g}}{1 \text{ mol}} = 2.74 \text{ g SO}_3$$

3. What mass of CO_2 gas would form if 25.00 g of carbon monoxide is heated?
 $2 \text{ CO} + \text{O}_2 \rightarrow 2 \text{ CO}_2$



$$25.0 \text{ g CO} \times \frac{1 \text{ mol}}{28.01 \text{ g}} \times \text{stoichiometric factor} \left(\frac{2 \text{ mol CO}_2}{2 \text{ mol CO}} \right) = 0.8926 \text{ mol CO}_2 \times \frac{44.01 \text{ g}}{1 \text{ mol}} = 39.3 \text{ g CO}_2$$

4. The reaction between 11.8 g of CH_4 and excess oxygen gas produces 11.8 g of CO_2 gas and some water. Determine the percent yield.



$$11.8 \text{ g CH}_4 \times \frac{1 \text{ mol}}{16.04 \text{ g}} \times \text{stoichiometric factor} \left(\frac{1 \text{ mol CO}_2}{1 \text{ mol CH}_4} \right) = 0.7356 \text{ mol CO}_2 \times \frac{44.01 \text{ g}}{1 \text{ mol}} = 32.38 \text{ g CO}_2$$

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