

# Classifying and Balancing Equations Multiple Choice

Name \_\_\_\_\_  
Date \_\_\_\_\_

Per \_\_\_\_\_

## PRACTICE TEST

- \_\_\_\_\_ 1. A chemical reaction is a process in which
  - a. products change into reactants
  - b. the law of conservation of mass applies
  - c. substances can change state
  - d. all of these
- \_\_\_\_\_ 2. During a chemical reaction,
  - a. new elements are produced
  - b. atoms are rearranged
  - c. atoms are destroyed
  - d. elements are destroyed
- \_\_\_\_\_ 3. An equation is balanced by
  - a. changing subscripts
  - b. adding coefficients
  - c. erasing elements as necessary
  - d. adding elements as necessary
- \_\_\_\_\_ 4. An atom's ability to undergo chemical reactions is determined by its
  - a. protons
  - b. innermost electrons
  - c. neutrons
  - d. valence electrons
- \_\_\_\_\_ 5. What are the reactants in the following chemical equation:  
$$\text{Zn} + \text{CuSO}_4 \longrightarrow \text{ZnSO}_4 + \text{Cu}$$
  - a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc
- \_\_\_\_\_ 6. What are the products in the above equation?
  - a. zinc and copper
  - b. zinc sulfate and copper
  - c. zinc and copper (II) sulfate
  - d. only zinc

## Short Answer

For questions 7-12, classify the reaction according to the type it is. Put that answer in the blank. Then add coefficients to balance the reaction when necessary.

- \_\_\_\_\_ 7.  $\text{Zn} + \text{H}_2\text{SO}_4 \longrightarrow \text{ZnSO}_4 + \text{H}_2$
- \_\_\_\_\_ 8.  $\text{H}_2\text{CO}_3 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- \_\_\_\_\_ 9.  $\text{CaCO}_3 + \text{HCl} \longrightarrow \text{CaCl}_2 + \text{H}_2\text{CO}_3$
- \_\_\_\_\_ 10.  $\text{AgNO}_3 + \text{Zn} \longrightarrow \text{Zn}(\text{NO}_3)_2 + \text{Ag}$
- \_\_\_\_\_ 11.  $\text{C}_3\text{H}_8 + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$
- \_\_\_\_\_ 12.  $\text{C}_2\text{H}_5\text{OH} + \text{O}_2 \longrightarrow \text{CO}_2 + \text{H}_2\text{O}$

Write a balanced equation for each of the following reactions:

13. Magnesium chloride is the product of a reaction between magnesium and chlorine.
14. Copper (II) hydroxide and potassium sulfate are produced when potassium hydroxide reacts with copper (II) sulfate.

## Unit 6 Practice Test - Key

- |      |                                                                           |                                                                                                     |
|------|---------------------------------------------------------------------------|-----------------------------------------------------------------------------------------------------|
| 1. b | 6. B                                                                      | 11. Combustion 1, 5, 3, 4                                                                           |
| 2. b | 7. Single replacement 1,1,1,1                                             | 12. Combustion 1,3,2,3                                                                              |
| 3. b | 8. Decomposition 1,1,1                                                    | 13. $\text{Mg} + \text{Cl}_2 \rightarrow \text{MgCl}_2$ 1,1,1                                       |
| 4. d | 9. Double replacement 1,2,1,1-<br>$\text{CaCO}_3$ is ppt (see Ref. Table) | 14. $\text{KOH} + \text{CuSO}_4 \rightarrow \text{Cu}(\text{OH})_2 + \text{K}_2\text{SO}_4$ 2,1,1,1 |
| 5. c | 10. Single replacement 2,1,1,2                                            |                                                                                                     |

# Modeling Chemistry Unit 7 Chemical Reactions Key

**Herbert DaCosta, Maohong Fan**



## **Modeling Chemistry Unit 7 Chemical Reactions Key:**

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**Green Chemistry** Mark Anthony Benvenuto, Steven Kosmas, 2022-08-22 This volume includes several perspectives on how to connect the United Nations Sustainable Development Goals with the 12 principles of green chemistry and green chemistry education

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**Principles of Modern Chemistry** David W. Oxtoby,H. Pat

Gillis,Laurie J. Butler,2016-01-01 Long considered the standard for honors and high level mainstream general chemistry courses PRINCIPLES OF MODERN CHEMISTRY continues to set the standard as the most modern rigorous and chemically and mathematically accurate text on the market This authoritative text features an atoms first approach and thoroughly revised chapters on Quantum Mechanics and Molecular Structure Chapter 6 Electrochemistry Chapter 17 and Molecular Spectroscopy and Photochemistry Chapter 20 In addition the text utilizes mathematically accurate and artistic atomic and molecular orbital art and is student friendly without compromising its rigor End of chapter study aids focus on only the most important key objectives equations and concepts making it easier for students to locate chapter content while applications to a wide range of disciplines such as biology chemical engineering biochemistry and medicine deepen students understanding of the relevance of chemistry beyond the classroom *Environmental Health Perspectives* ,1993 *Selected Water*

*Resources Abstracts* ,1991 **A Two-dimensional Numerical Simulation of a Supersonic, Chemically Reacting Mixing Layer** John Philip Drummond,1988 Rate Constant Calculation for Thermal Reactions Herbert DaCosta, Maohong Fan,2011-12-28 Providing an overview of the latest computational approaches to estimate rate constants for thermal reactions this book addresses the theories behind various first principle and approximation methods that have emerged in the last twenty years with validation examples It presents in depth applications of those theories to a wide range of basic and applied research areas When doing modeling and simulation of chemical reactions as in many other cases one often has to compromise between higher accuracy higher precision approaches which are usually time consuming and approximate lower precision approaches which often has the advantage of speed in providing results This book covers both approaches It is augmented by a wide range of applications of the above methods to fuel combustion unimolecular and bimolecular reactions isomerization polymerization and to emission control of nitrogen oxides An excellent resource for academics and industry members in physical chemistry chemical engineering and related fields *CHEMISTRY* SILBERBERG,2003

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Green Chemical Engineering, Volume 12 ,2018-12-10 Green chemistry and chemical engineering belong together and this twelfth volume in the successful Handbook of Green Chemistry series represents the perfect one stop reference on the topic Written by an international team of specialists with each section edited by international leading experts this book provides first hand insights into the field covering chemical engineering process design innovations in unit operations and manufacturing biorefining and much more besides An indispensable source for every chemical engineer in industry and academia

*NBS Special Publication* ,1979      Trace Gas Exchange in Forest Ecosystems R. Gasche,H. Papen,H. Rennenberg,2013-03-14 This volume summarizes the current knowledge on the exchange of trace gases between forests and the atmosphere with the restriction that exclusively carbon and nitrogen compounds are included For this purpose the volume brings together and interconnects knowledge from different disciplines of biological and atmospheric sciences It covers microbial and plant processes involved in the production and consumption of these trace gases the exchange processes between forest soils and vegetation on the one hand and the atmosphere on the other hand the fate of the trace gases exchanged inside the atmosphere as well as environmental influences on the exchange of trace gases between forest ecosystems and the atmosphere With this interdisciplinary approach the volume provides the background for an evaluation of the exchange of trace gases between forest ecosystems and the atmosphere and man made disturbances of this exchange

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