

Chemistry Unit 8 Worksheet 4

Samples of Every Kind of Problem

On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.

1. Calculate the number of moles of potassium chlorate, KClO_3 (s), that must decompose to produce potassium chloride, KCl (s), and 1.8 moles of oxygen gas.

| | | | | |
|-----------|---|------|------|-------------------------------------|
| Equation: | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ | | | 1.2 moles KClO_3 needed |
| Before: | 1.2 | 0 | 0 | |
| Change: | -1.2 | +1.2 | +1.8 | |
| After: | 0 | 1.2 | 1.8 | |

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

| | | | | | |
|-----------|---|----------|---------|-------|------------------------------------|
| Equation: | $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ | | | | 0.200 moles HCl needed |
| Before: | 0.100 | ? | 0 | 0 | |
| Change: | -0.100 | -0.200 + | 0.100 + | 0.100 | |
| After: | 0 | ? | 0.100 | 0.100 | |

3. Ethane, C_2H_6 reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

| | | | | |
|-----------|--|-------|-------|-------|
| Equation: | $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ | | | |
| Before: | 2.20 | xs | 0 | 0 |
| Change: | -2.20 | -7.70 | +4.40 | +6.60 |
| After: | 0 | xs | 4.40 | 6.60 |

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

| | | | | |
|-----------|---|---------|---------|---------|
| Equation: | $\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$ | | | |
| Before: | 0.0040 | XS | 0 | 0 |
| Change: | -0.0040 | -0.0080 | +0.0040 | +0.0080 |
| After: | 0 | XS | 0.0040 | 0.0080 |

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

Modeling Chemistry Unit 4 4 Answers

ML Morrison



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Table of Contents Modeling Chemistry Unit 4 4 Answers

1. Understanding the eBook Modeling Chemistry Unit 4 4 Answers
 - The Rise of Digital Reading Modeling Chemistry Unit 4 4 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Modeling Chemistry Unit 4 4 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Modeling Chemistry Unit 4 4 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Modeling Chemistry Unit 4 4 Answers
 - Personalized Recommendations
 - Modeling Chemistry Unit 4 4 Answers User Reviews and Ratings
 - Modeling Chemistry Unit 4 4 Answers and Bestseller Lists
5. Accessing Modeling Chemistry Unit 4 4 Answers Free and Paid eBooks

- Modeling Chemistry Unit 4 4 Answers Public Domain eBooks
 - Modeling Chemistry Unit 4 4 Answers eBook Subscription Services
 - Modeling Chemistry Unit 4 4 Answers Budget-Friendly Options
6. Navigating Modeling Chemistry Unit 4 4 Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Modeling Chemistry Unit 4 4 Answers Compatibility with Devices
 - Modeling Chemistry Unit 4 4 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Modeling Chemistry Unit 4 4 Answers
 - Highlighting and Note-Taking Modeling Chemistry Unit 4 4 Answers
 - Interactive Elements Modeling Chemistry Unit 4 4 Answers
8. Staying Engaged with Modeling Chemistry Unit 4 4 Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Modeling Chemistry Unit 4 4 Answers
9. Balancing eBooks and Physical Books Modeling Chemistry Unit 4 4 Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Modeling Chemistry Unit 4 4 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Modeling Chemistry Unit 4 4 Answers
- Setting Reading Goals Modeling Chemistry Unit 4 4 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Modeling Chemistry Unit 4 4 Answers
- Fact-Checking eBook Content of Modeling Chemistry Unit 4 4 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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