

Stoichiometry Worksheet 2: Percent Yield

What aspects of the procedure do you think:

- affected the theoretical or actual yield negatively?
- substantially affected your yield in a positive or negative way? explain how these things can be made up.
- affected your yield in a positive or negative way? explain why these things can be made up.

10. Using the theoretical yield and the actual yield, calculate the percent yield of the reaction. Show the calculation and answer. What does the percent yield tell you about the reaction? (Percent yield is the ratio of actual yield to theoretical yield.)

$$\text{Theoretical Yield} = 1.00 \text{ g} \times \frac{100\%}{100\%} = 1.00 \text{ g}$$

$$\text{Actual Yield} = 0.80 \text{ g}$$

$$\text{Percent Yield} = \frac{0.80 \text{ g}}{1.00 \text{ g}} \times 100\% = 80\%$$

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$$\begin{aligned} \text{Theoretical Yield} &= 1.00 \text{ g} \times \frac{100\%}{100\%} = 1.00 \text{ g} \\ \text{Actual Yield} &= 0.80 \text{ g} \\ \text{Percent Yield} &= \frac{0.80 \text{ g}}{1.00 \text{ g}} \times 100\% = 80\% \end{aligned}$$

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Modeling Chemistry U8 Ws 1 V2

Jessica J Manson



Modeling Chemistry U8 Ws 1 V2 :

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