

## Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{1 \text{ mole}}{63.5 \text{ g}} = 5.2 \times 10^{-2} \text{ mol or } 0.052 \text{ mol}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{1 \text{ mole}}{55.8 \text{ g}} = 7.92 \times 10^{-2} \text{ mol or } 0.0792 \text{ mol}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$0.050 \text{ g} \times \frac{1 \text{ mole}}{18.0 \text{ g}} = 2.78 \times 10^{-3} \text{ mol} \rightarrow 2.8 \times 10^{-3} \text{ mol}$$

4. What mass of water would you need to have 15.0 moles of H<sub>2</sub>O?

$$15.0 \text{ moles} \times \frac{18.0 \text{ g}}{1 \text{ mole}} = 2.70 \times 10^2 \text{ g}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride is this?

$$737 \text{ g} \times \frac{1 \text{ mole}}{58.5 \text{ g}} = 12.6 \text{ mol}$$

Na	23.0
Cl	35.5
	<b>58.5 g</b>

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate). How many grams should the chef mass out?

$$0.050 \text{ moles} \times \frac{84.0 \text{ g}}{1 \text{ mole}} = 4.2 \text{ g}$$

Na	23.0
H	1.0
C	12.0
O	48.0
	<b>84.0</b>

7. Rust is iron(III) oxide. The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron(III) oxide are contained in 2.50 kg of rust?

$$2.50 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mole}}{159.6 \text{ g}} = 15.7 \text{ mol Fe}_2\text{O}_3$$

Fe	111.6
O	48.0
	<b>159.6</b>

# Modeling Chemistry Unit Ws 2 Answers

**Mike Jess**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Mike Jess' text.

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web limiting reactant theoretical and percentage yields key  $4 \text{KO}_2 + 2 \text{H}_2\text{O} \rightarrow 4 \text{KOH} + 3 \text{O}_2$  g if a reaction vessel contains 0 mol  $\text{KO}_2$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant

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 $3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$

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web start with 8.3 mole H<sub>2</sub>S and multiply it with 2 mole H<sub>2</sub>O over 2 mole H<sub>2</sub>S from the equation  
 $8.3 \text{ mole H}_2\text{S} \times \frac{2 \text{ mole H}_2\text{O}}{2 \text{ mole H}_2\text{S}} \times 18 \text{ g H}_2\text{O} = 149.4 \text{ g H}_2\text{O}$   
 149.4 g H<sub>2</sub>O is your theoretical yield so percent yield =  $\frac{137.1 \text{ g H}_2\text{O}}{149.4 \text{ g H}_2\text{O}} \times 100 = 91.77\%$   
 problems 6 & 8 are similar to problem 5

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web 1 csun.edu/hcchm001/introchemhandouts.html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of NH<sub>3</sub> can be produced theoretically from the reaction of 5.0 g of N  
 what is the limiting reagent if 8.52 g are actually formed what is the

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web limiting reagents problem one 1 for the balanced equation  
 $\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$   
 if the reaction of 21.3 grams of O<sub>2</sub> produces a 47.3% yield how many grams of CO would be produced  
 step one figure out how many moles of the limiting reagent you have  
 $21.3 \text{ g O}_2 \times \frac{1 \text{ mole}}{32 \text{ g}} = 0.666 \text{ moles O}_2$   
 $0.666 \text{ moles O}_2 \times \frac{2 \text{ moles CO}}{3 \text{ moles O}_2} = 0.444 \text{ moles CO}$   
 $0.444 \text{ moles CO} \times 28 \text{ g/mole} = 12.4 \text{ g CO}$

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web percent yield is a measure of the efficiency of the experimental design  
 yield efficiency =  $\frac{\text{mass of product obtained}}{\text{calculated mass of the product expected}} \times 100$   
 in this reaction a yield of 80% is anticipated

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web a write a balanced equation for the reaction  
 $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$   
 4 Al + 3 O<sub>2</sub> → 2 Al<sub>2</sub>O<sub>3</sub>  
 b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub>  
 $160.0 \text{ g Al} \times \frac{1 \text{ mole Al}}{27 \text{ g Al}} \times \frac{2 \text{ moles Al}_2\text{O}_3}{4 \text{ moles Al}} \times 101.96 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$   
 c determine the percent yield  
 $\frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$   
 theoretical = 302.3 g

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield = 86.68% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless  
 10 nitrogen is the limiting reagent  
 11 yes methane is the limiting reagent  
 12 C is the limiting reagent  
 4.33 g of H<sub>2</sub> are left over

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remains when the reaction is complete

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80 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2$  x 133.8 g  $\text{I}_2\text{O}_5$  1 mol  $\text{I}_2\text{O}_5$  28 g CO 1 mol CO

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage  $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$  label 3 7 3

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