

Chapter 8: Particle Physics

Outline the results of the gold leaf experiments

1. Alpha particles fired at gold metal foil
2. In a vacuum
3. Most pass straight through/undeflected;
4. Some deflected;
5. Very few reflected/deflected through large angles;
6. Atom mainly empty space;
7. Most did not get near enough to any matter to be deflected;
8. Only a few particles came close enough to charge to be affected;
9. Nucleus has mass much greater than alpha particle mass to cause deflection;
10. Nucleus contains most of the mass;
11. Nucleus very small/tiny;
12. Nucleus charged/positive;
13. Alpha particles are deflected by a charged nucleus that has a very small diameter compared to that of the atom rather than the charge being distributed throughout the atom;
14. most of the mass of the atom is concentrated in the nucleus rather than distributed throughout the atom;

Explain how the observations from the alpha particle scattering experiment led to a new model of the atom

1. Most alpha particles pass through undeflected OR some deflected through a small angle;
2. A very small number are deflected through an angle greater than 90 degrees;
3. This suggests that the alpha particles are deflected by a charged nucleus that has a very small diameter compared to that of the atom rather than the charge being distributed throughout the atom;
4. and that most of the mass of the atom is concentrated in the nucleus rather than distributed throughout the atom;

Role of E and B fields in particle detector

1. Electric field accelerates particles
2. Direction of deflection/force indicates sign of charge;
3. $a = EQ/m$;
4. Magnetic fields provide a centripetal force*;
5. Direction of force/curvature indicates sign of charge;
6. Momentum from radius;
7. $r = p/BQ$;

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Michael Brown

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Table of Contents Jan 2015 Physics Mark Scheme

1. Understanding the eBook Jan 2015 Physics Mark Scheme
 - The Rise of Digital Reading Jan 2015 Physics Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying Jan 2015 Physics Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jan 2015 Physics Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from Jan 2015 Physics Mark Scheme
 - Personalized Recommendations
 - Jan 2015 Physics Mark Scheme User Reviews and Ratings

- Jan 2015 Physics Mark Scheme and Bestseller Lists
- 5. Accessing Jan 2015 Physics Mark Scheme Free and Paid eBooks
 - Jan 2015 Physics Mark Scheme Public Domain eBooks
 - Jan 2015 Physics Mark Scheme eBook Subscription Services
 - Jan 2015 Physics Mark Scheme Budget-Friendly Options
- 6. Navigating Jan 2015 Physics Mark Scheme eBook Formats
 - ePub, PDF, MOBI, and More
 - Jan 2015 Physics Mark Scheme Compatibility with Devices
 - Jan 2015 Physics Mark Scheme Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jan 2015 Physics Mark Scheme
 - Highlighting and Note-Taking Jan 2015 Physics Mark Scheme
 - Interactive Elements Jan 2015 Physics Mark Scheme
- 8. Staying Engaged with Jan 2015 Physics Mark Scheme
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jan 2015 Physics Mark Scheme
- 9. Balancing eBooks and Physical Books Jan 2015 Physics Mark Scheme
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Jan 2015 Physics Mark Scheme
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Jan 2015 Physics Mark Scheme
 - Setting Reading Goals Jan 2015 Physics Mark Scheme
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Jan 2015 Physics Mark Scheme
 - Fact-Checking eBook Content of Jan 2015 Physics Mark Scheme
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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