

TOPIC 8 : HEAT & TEMPERATURE

Specific Heat Capacity :-

$$\Delta E = mc\Delta\theta$$

where; c = specific heat capacity

m = mass

$\Delta\theta$ = change in temperature

Specific Latent Heat :-

$$\Delta E = L\Delta m$$

L = specific latent heat

Kinetic Energy of molecule in a gaseous sample

$$\frac{1}{2}m\langle c^2 \rangle = \frac{3}{2}kT$$

where; k = Boltzmann constant
 $1.38 \times 10^{-23} \text{ J K}^{-1}$

T = temperature in Kelvin

GIAS LAWS

• $V \propto T$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

• $p \propto T$

• $p \propto \frac{1}{V}$ $p_1V_1 = p_2V_2$

where, p = pressure

T = temp.

V = vol.

Gras laws ultimate equation

$$pV = NkT$$

where; N = number of molecules in gas

k = Boltzmann const.

T = temp. in Kelvin

Equation of state for ideal gas

$$pV = nRT$$

$$pV = \frac{1}{3}mN\langle c^2 \rangle$$

where; n = number of moles of gas

R = universal gas constant = $8.31 \text{ J kg}^{-1} \text{ mol}^{-1}$

Jan 2015 Ial Unit 5 Physics Paper

Robert R. Redfield



Jan 2015 Ial Unit 5 Physics Paper:

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Table of Contents Jan 2015 Ial Unit 5 Physics Paper

1. Understanding the eBook Jan 2015 Ial Unit 5 Physics Paper
 - The Rise of Digital Reading Jan 2015 Ial Unit 5 Physics Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Jan 2015 Ial Unit 5 Physics Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Jan 2015 Ial Unit 5 Physics Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Jan 2015 Ial Unit 5 Physics Paper
 - Personalized Recommendations
 - Jan 2015 Ial Unit 5 Physics Paper User Reviews and Ratings
 - Jan 2015 Ial Unit 5 Physics Paper and Bestseller Lists

5. Accessing Jan 2015 Ial Unit 5 Physics Paper Free and Paid eBooks
 - Jan 2015 Ial Unit 5 Physics Paper Public Domain eBooks
 - Jan 2015 Ial Unit 5 Physics Paper eBook Subscription Services
 - Jan 2015 Ial Unit 5 Physics Paper Budget-Friendly Options
6. Navigating Jan 2015 Ial Unit 5 Physics Paper eBook Formats
 - ePub, PDF, MOBI, and More
 - Jan 2015 Ial Unit 5 Physics Paper Compatibility with Devices
 - Jan 2015 Ial Unit 5 Physics Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Jan 2015 Ial Unit 5 Physics Paper
 - Highlighting and Note-Taking Jan 2015 Ial Unit 5 Physics Paper
 - Interactive Elements Jan 2015 Ial Unit 5 Physics Paper
8. Staying Engaged with Jan 2015 Ial Unit 5 Physics Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Jan 2015 Ial Unit 5 Physics Paper
9. Balancing eBooks and Physical Books Jan 2015 Ial Unit 5 Physics Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Jan 2015 Ial Unit 5 Physics Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Jan 2015 Ial Unit 5 Physics Paper
 - Setting Reading Goals Jan 2015 Ial Unit 5 Physics Paper
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Jan 2015 Ial Unit 5 Physics Paper
 - Fact-Checking eBook Content of Jan 2015 Ial Unit 5 Physics Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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