(2)

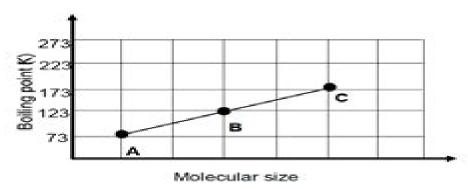
(1)

(3) [13]

#### QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The **letters A**, **B** and **C** represent the compounds CH<sub>4</sub>, C<sub>4</sub>H<sub>6</sub> and C<sub>3</sub>H<sub>8</sub> respectively.



- 2.1.1 Define the term boiling point.
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (A, B or C) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph.
  (2)
- 2.2 Consider the two molecules in the table below.

NAME OFSUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT
Ammonia	NH <sub>3</sub>	17	- 33
Phosphine	PH <sub>3</sub>	×	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES.

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# **Table of Contents Grade 1 previous Question Papers Physical Science**

- 1. Understanding the eBook Grade 1 previous Question Papers Physical Science
  - The Rise of Digital Reading Grade 1 previous Question Papers Physical Science
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Grade 1 previous Question Papers Physical Science
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade 1 previous Question Papers Physical Science
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Grade 1 previous Question Papers Physical Science
  - Personalized Recommendations
  - Grade 1 previous Question Papers Physical Science User Reviews and Ratings
  - Grade 1 previous Question Papers Physical Science and Bestseller Lists

- 5. Accessing Grade 1 previous Question Papers Physical Science Free and Paid eBooks
  - Grade 1 previous Question Papers Physical Science Public Domain eBooks
  - Grade 1 previous Question Papers Physical Science eBook Subscription Services
  - Grade 1 previous Question Papers Physical Science Budget-Friendly Options
- 6. Navigating Grade 1 previous Question Papers Physical Science eBook Formats
  - ∘ ePub, PDF, MOBI, and More
  - Grade 1 previous Question Papers Physical Science Compatibility with Devices
  - Grade 1 previous Question Papers Physical Science Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grade 1 previous Question Papers Physical Science
  - Highlighting and Note-Taking Grade 1 previous Question Papers Physical Science
  - Interactive Elements Grade 1 previous Question Papers Physical Science
- 8. Staying Engaged with Grade 1 previous Question Papers Physical Science
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Grade 1previous Question Papers Physical Science
- 9. Balancing eBooks and Physical Books Grade 1 previous Question Papers Physical Science
  - ∘ Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Grade 1 previous Question Papers Physical Science
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Grade 1 previous Question Papers Physical Science
  - Setting Reading Goals Grade 1 previous Question Papers Physical Science
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade 1 previous Question Papers Physical Science
  - Fact-Checking eBook Content of Grade 1 previous Question Papers Physical Science
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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