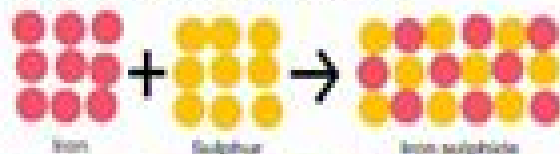


Topic 2: Chemical Change (P 114-129)

Unit 1: Physical & Chemical Change

- Physical change = change in matter that can be reversed.
- 2 substances that are mixed and can be separated by physical means.
- E.g. solid water (ice) is arranged in crystal lattice.
 - Heat breaks the lattice - transforming it into liquid water and gas (water vapor).
 - Atoms in water remain bonded due to strong forces.
- Chemical change = changing something so that it can't be changed back to its original state.
- E.g. heating iron and sulphur together creates a brown, sticky substance called iron sulphide (FeS).
 - This new substance has different properties from the old one (FeS isn't magnetic).
 - Heating sulphur with iron causes atoms to break apart and rejoin in a strong chemical bond - new lattice formed.



- Chemical changes require more energy than physical ones.

Physical Change	Chemical Change
Easy to reverse.	Difficult to reverse.
No new chemical substances form.	New chemical substances form.
Require small amounts of energy.	Require large amounts of energy.

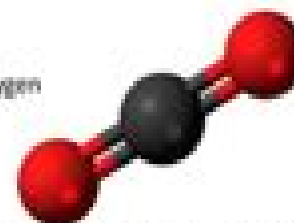
Unit 2: Conservation of Atoms During Chemical Change

Law of Conservation of Matter

- No. of atoms doesn't change when a chemical reaction occurs.
- No. molecules may change.
- Law of conservation of matter = mass of reactants and mass of products stays the same.

Law of Constant Proportions

- Law of constant proportions = a chemical compound always contains the same elements in fixed proportions by mass.
- E.g. CO_2 :
 - Relative atomic mass of carbon = 12.
 - Relative atomic mass of oxygen = 16 (2 oxygen atoms required in CO_2).
 - Total relative atomic mass of CO_2 is $12 + 16 + 16 = 44$.



- Ratio of carbon to oxygen remains constant - any mass of CO_2 will be $\frac{12}{44}$ carbon and $\frac{32}{44}$ oxygen.
- Ratio of mass of carbon to oxygen is always 12:32 or 3:8 in any amount of CO_2 , irrespective of how it is formed.

Unit 3: Representing Chemical Change

Balanced Chemical Equations

- Chemical change is represented by chemical equations.
- Substances on the left of the arrow = reactants.
- Substances on the right of the arrow = products.

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Subject Offerings and Enrollments, Grades 9-12 Diane Bochner Gertler, 1965

Resources in Education , 1991-04

Progreso de la Instrucción Pública en Los Estados Unidos de América United States. Office of Education, 1962

Science Education in Countries Along the Belt & Road Ronghuai Huang, Bing Xin, Ahmed Tlili, Feng Yang, Xiangling Zhang, Lixin Zhu, Mohamed Jemni, 2022-01-18 This book aims to highlight science education in countries along the Belt and Road It consists of 30 chapters divided into three main parts namely Arab and African countries Asian countries and European countries We invited science education experts from 29 Belt and Road countries to introduce the current status of science education in their countries and the new requirements with the rapid evolution of Information Technology The major contributions of this book include 1 Provide the current status of science education in countries along the Belt and Road as well as the requirement for developing and improving science education in these countries 2 Discuss new insights of science education in future years 3 Inspire stakeholders to take effective initiatives to develop science education in countries along the Belt and Road

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