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CHAPTER 12**STUDY GUIDE FOR CONTENT MASTERY****Section 12.3 Limiting Reactants**

In your textbook, read about why reactions stop and how to determine the limiting reactant.

Study the diagram showing a chemical reaction and the chemical equation that represents the reaction. Then complete the table. Show your calculations for questions 15–17 in the space below the table.



The molar masses of O_2 , NO , and NO_2 are 32.00 g/mol, 30.01 g/mol, and 46.01 g/mol, respectively.

Amount of O_2	Amount of NO	Amount of NO_2	Limiting Reactant	Amount and Mass of Excess Reactant
1 molecule	2 molecules	2 molecules	none	none
4 molecules	4 molecules	4 molecules	NO	2 molecules O_2
2 molecules	4 molecules	4 molecules	O_2	2 molecules NO
1.00 mol	2.00 mol	2.00 mol	none	none
4.00 mol	4.00 mol	4.00 mol	NO	2.00 mol O_2
0.00 mol	0.00 mol	0.00 mol	NO	0.00 mol O_2
1.00 mol	0.00 mol	0.00 mol	O_2	0.00 mol NO
0.000 mol	0.000 mol	0.000 mol	NO	0.000 mol O_2
32.00 g	60.02 g	92.02 g	none	none
64.00 g	60.02 g	46.01 g	O_2	30.01 g NO
64.00 g	30.01 g	28.76 g	O_2	1.24 g NO

balanced equation mole ratio = 2 mol NO /1 mol O_2

$10.00 \text{ g } \text{O}_2 \times 1 \text{ mol } \text{O}_2 / 32.00 \text{ g } \text{O}_2 = 0.3125 \text{ mol } \text{O}_2$

$20.00 \text{ g } \text{NO} \times 1 \text{ mol } \text{NO} / 30.01 \text{ g } \text{NO} = 0.6664 \text{ mol } \text{NO}$

actual mole ratio = $0.6664 \text{ mol } \text{NO} / 0.3125 \text{ mol } \text{O}_2 = 2.132 \text{ mol } \text{NO} / 1.000 \text{ mol } \text{O}_2$

Because the actual mole ratio of $\text{NO}:\text{O}_2$ is larger than the balanced equation mole ratio of $\text{NO}:\text{O}_2$, there is an excess of NO ; O_2 is the limiting reactant.

Mass of NO used = $0.3125 \text{ mol } \text{O}_2 \times 2 \text{ mol } \text{NO} / 1 \text{ mol } \text{O}_2 = 0.6250 \text{ mol } \text{NO}$

$0.6250 \text{ mol } \text{NO} \times 30.01 \text{ g } \text{NO} / 1 \text{ mol } \text{NO} = 18.76 \text{ g } \text{NO}$

Mass of NO_2 produced = $0.6250 \text{ mol } \text{NO}_2 \times 46.01 \text{ g } \text{NO}_2 / 1 \text{ mol } \text{NO}_2 = 28.76 \text{ g } \text{NO}_2$

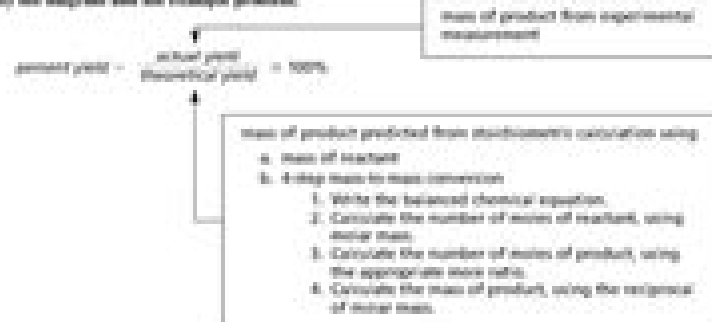
Excess $\text{NO} = 20.00 \text{ g } \text{NO} - 18.76 \text{ g } \text{NO} = 1.24 \text{ g } \text{NO}$

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CHAPTER 13**STUDY GUIDE FOR CONTENT MASTERY****Section 13.4 Percent Yield**

In your textbook, read about the yields of products.

Study the diagram and the example problem.



Example Problem: The following chemical equation represents the production of palladium, a valuable metal in the manufacturing of some semiconductor devices.



In one experiment, the reaction yielded 7.42 g of the metal from a 7.00-g sample of palladium. Determine the percent yield of this reaction. The molar masses of Ga and Ga_2O_3 are 69.72 g/mol and 157.44 g/mol, respectively.

Use the information in the diagram and example problem to evaluate each value or expression below. If the value or expression is correct, write correct. If it is incorrect, write the correct value or expression.

- actual yield unknown 7.42 g Ga_2O_3
- mass of reactant: 7.00 g Ga correct
- number of moles of reactant: $7.00 \text{ g } \text{Ga} \times \frac{69.72 \text{ g } \text{Ga}}{1 \text{ mol } \text{Ga}} = 1 \text{ mol } \text{Ga}$ 7.00 g $\text{Ga} \times 1 \text{ mol } \text{Ga} / 69.72 \text{ g } \text{Ga}$
- number of moles of product: $0.100 \text{ mol } \text{Ga} \times \frac{2 \text{ mol } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } \text{Ga}} = 0.200 \text{ mol } \text{Ga}_2\text{O}_3$ 0.100 mol $\text{Ga} \times \frac{2 \text{ mol } \text{Ga}_2\text{O}_3}{1 \text{ mol } \text{Ga}}$
- theoretical yield: $0.200 \text{ mol } (\text{Ga}_2\text{O}_3) \times \frac{157.44 \text{ g } (\text{Ga}_2\text{O}_3)}{1 \text{ mol } (\text{Ga}_2\text{O}_3)} = 31.49 \text{ g } \text{Ga}_2\text{O}_3$ correct
- percent yield: $\frac{7.42 \text{ g } (\text{Ga}_2\text{O}_3)}{31.49 \text{ g } (\text{Ga}_2\text{O}_3)} \times 100 = 23.56\%$ 7.42 g $\text{Ga}_2\text{O}_3 / 31.49 \text{ g } \text{Ga}_2\text{O}_3 \times 100$

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Ying Liu



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1. Understanding the eBook Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - The Rise of Digital Reading Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Mcgraw Hill Chemistry Chap 26 Assessment Answers

- Personalized Recommendations
 - Mcgraw Hill Chemistry Chap 26 Assessment Answers User Reviews and Ratings
 - Mcgraw Hill Chemistry Chap 26 Assessment Answers and Bestseller Lists
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 - Highlighting and Note-Taking Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - Interactive Elements Mcgraw Hill Chemistry Chap 26 Assessment Answers
 8. Staying Engaged with Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Mcgraw Hill Chemistry Chap 26 Assessment Answers
 9. Balancing eBooks and Physical Books Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Mcgraw Hill Chemistry Chap 26 Assessment Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - Setting Reading Goals Mcgraw Hill Chemistry Chap 26 Assessment Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Mcgraw Hill Chemistry Chap 26 Assessment Answers

- Fact-Checking eBook Content of Mcgraw Hill Chemistry Chap 26 Assessment Answers
- Distinguishing Credible Sources

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- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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