

PHYSICAL: Kinetic Molecular Theory

Learning Objectives

- At least a theoretical understanding of how gases are represented by the kinetic theory and an understanding of the limitations of the kinetic molecular theory
- An ability to quantitatively calculate pressure, volume, and temperature
- An ability to qualitatively explain the behavior of gases in terms of the kinetic theory

Module 1: Representation of Gases in Different Phases

The following figure illustrates the representation of a gas in three different phases: solid, liquid, and gas.

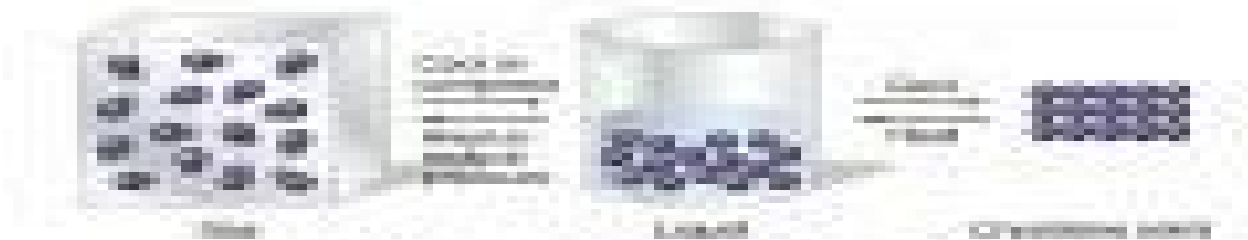


Figure 1: Representation of a gas in three different phases: solid, liquid, and gas.

Understanding the representation of gases in different phases is essential for understanding the behavior of gases in different environments.

Key Concepts

- **Pressure:** The force exerted by a gas on the walls of its container. It is measured in Pascals (Pa) or Newtons per square meter (N/m²). The pressure of a gas is determined by the number of gas molecules and the volume of the container.

Figure 2: Representation of a gas in three different phases: solid, liquid, and gas.

	SOLID	LIQUID	GAS
Particle arrangement	Particles are packed closely together in a regular, repeating pattern. The particles are held together by strong forces.	Particles are packed closely together in a disordered, irregular pattern. The particles are held together by moderate forces.	Particles are spread out in a disordered, irregular pattern. The particles are held together by weak forces.
Particle movement	Particles vibrate in place. They have a fixed position and can only move a small distance from their equilibrium position.	Particles move around in a disordered, irregular pattern. They have a fixed position and can move a small distance from their equilibrium position.	Particles move around in a disordered, irregular pattern. They have a fixed position and can move a large distance from their equilibrium position.

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Ian Pickup



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