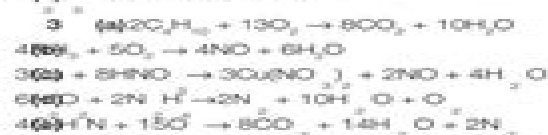
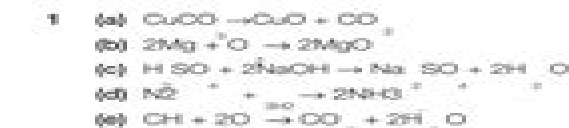


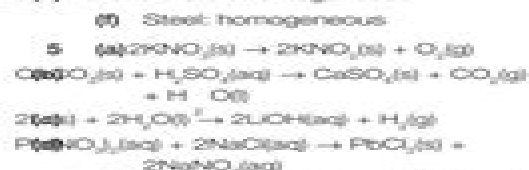
Answers

Chapter 1

Exercises



4 (a) Sand and water: heterogeneous
 (b) Sugar and water: homogeneous
 (c) Oil and iron filings: heterogeneous
 (d) Endol and water: homogeneous



In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)

7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.

8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.

9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.

10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E–D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.

11 B

12 AW 01_EX_12

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1. Understanding the eBook Lesson Pearson Chemistry Answers
 - The Rise of Digital Reading Lesson Pearson Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson Pearson Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson Pearson Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson Pearson Chemistry Answers
 - Personalized Recommendations
 - Lesson Pearson Chemistry Answers User Reviews and Ratings
 - Lesson Pearson Chemistry Answers and Bestseller Lists
5. Accessing Lesson Pearson Chemistry Answers Free and Paid eBooks
 - Lesson Pearson Chemistry Answers Public Domain eBooks
 - Lesson Pearson Chemistry Answers eBook Subscription Services
 - Lesson Pearson Chemistry Answers Budget-Friendly Options

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 - ePub, PDF, MOBI, and More
 - Lesson Pearson Chemistry Answers Compatibility with Devices
 - Lesson Pearson Chemistry Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson Pearson Chemistry Answers
 - Highlighting and Note-Taking Lesson Pearson Chemistry Answers
 - Interactive Elements Lesson Pearson Chemistry Answers
8. Staying Engaged with Lesson Pearson Chemistry Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson Pearson Chemistry Answers
9. Balancing eBooks and Physical Books Lesson Pearson Chemistry Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson Pearson Chemistry Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Lesson Pearson Chemistry Answers
 - Setting Reading Goals Lesson Pearson Chemistry Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Lesson Pearson Chemistry Answers
 - Fact-Checking eBook Content of Lesson Pearson Chemistry Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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