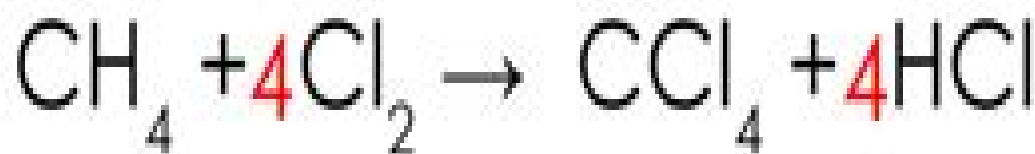


# Balancing Chemical Equations Practice

**Step 1** – Take Inventory of the elements and atoms on the product and reactant side.

**Step 2** – Is it balanced?

**Step 3** – If unbalanced, change coefficients until it's balanced.



$$\text{C} - 1 = \text{C} - 1$$

$$\text{H} - 4 \neq \text{H} - \cancel{1} 4$$

$$\text{Cl} - \cancel{2} 8 \neq \text{Cl} - \cancel{5} 8$$

**Balanced!**

# How To Balance Chemical Equations Study Guide

**Marcia L. Gillette, Ira Blei, Wendy  
Gloffke, George Odian**



## How To Balance Chemical Equations Study Guide:

*Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane  $\text{CH}_4$  from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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