

7 Ionic and Metallic Bonding

Big idea BONDING AND INTERACTIONS

7.1 Ions

For students using the Foundation edition, assign problems 1, 3–5, 7–12, 14, 15, 18–20.

Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged.

Lesson Summary

Valence Electrons Valence electrons are the electrons in the outermost occupied energy level and are involved in ion formation.

- ▶ For a representative element, the group number equals the number of valence electrons the atom contains.
- ▶ An electron dot structure shows the symbol of the element and its valence electrons.
- ▶ Atoms tend to gain or lose the number of electrons that will provide the atom with a noble gas electron configuration.

Formation of Cations Cations are positively charged ions formed when an atom loses one or more valence electrons.

- ▶ Atoms and the cations formed from them have different properties.
- ▶ Elements in Group 1 form cations with a charge of $1+$, and those in Group 2 form cations with a charge of $2+$.
- ▶ Many transition metals form more than one cation and do not follow the octet rule.

Formation of Anions Anions are negatively charged ions formed when an atom gains one or more valence electrons.

- ▶ Commonly, the name of an anion ends in *-ide*.
- ▶ Anions form from nonmetallic elements.
- ▶ The anions formed from halogens are known as halides.

After reading Lesson 7.1, answer the following questions.

Valence Electrons

1. What are valence electrons?

Valence electrons are the electrons in the highest occupied energy level of an element's atoms.

2. The valence electrons largely determine the *chemical properties* of an element and are usually the only electrons used in *chemical bonds*.

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Table of Contents Lesson 7 Reading And Study Workbook Answers Pearson

1. Understanding the eBook Lesson 7 Reading And Study Workbook Answers Pearson
 - The Rise of Digital Reading Lesson 7 Reading And Study Workbook Answers Pearson
 - Advantages of eBooks Over Traditional Books
2. Identifying Lesson 7 Reading And Study Workbook Answers Pearson
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Lesson 7 Reading And Study Workbook Answers Pearson
 - User-Friendly Interface
4. Exploring eBook Recommendations from Lesson 7 Reading And Study Workbook Answers Pearson
 - Personalized Recommendations
 - Lesson 7 Reading And Study Workbook Answers Pearson User Reviews and Ratings

- Lesson 7 Reading And Study Workbook Answers Pearson and Bestseller Lists
- 5. Accessing Lesson 7 Reading And Study Workbook Answers Pearson Free and Paid eBooks
 - Lesson 7 Reading And Study Workbook Answers Pearson Public Domain eBooks
 - Lesson 7 Reading And Study Workbook Answers Pearson eBook Subscription Services
 - Lesson 7 Reading And Study Workbook Answers Pearson Budget-Friendly Options
- 6. Navigating Lesson 7 Reading And Study Workbook Answers Pearson eBook Formats
 - ePub, PDF, MOBI, and More
 - Lesson 7 Reading And Study Workbook Answers Pearson Compatibility with Devices
 - Lesson 7 Reading And Study Workbook Answers Pearson Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Lesson 7 Reading And Study Workbook Answers Pearson
 - Highlighting and Note-Taking Lesson 7 Reading And Study Workbook Answers Pearson
 - Interactive Elements Lesson 7 Reading And Study Workbook Answers Pearson
- 8. Staying Engaged with Lesson 7 Reading And Study Workbook Answers Pearson
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Lesson 7 Reading And Study Workbook Answers Pearson
- 9. Balancing eBooks and Physical Books Lesson 7 Reading And Study Workbook Answers Pearson
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Lesson 7 Reading And Study Workbook Answers Pearson
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Lesson 7 Reading And Study Workbook Answers Pearson
 - Setting Reading Goals Lesson 7 Reading And Study Workbook Answers Pearson
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Lesson 7 Reading And Study Workbook Answers Pearson
 - Fact-Checking eBook Content of Lesson 7 Reading And Study Workbook Answers Pearson
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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