

- b) Calculate the density of NaCl, if edge length of NaCl unit cell is 564 pm. [Molar mass of NaCl = 58.5 g/mol]. (2)  
[March 2013]
- 11) Unit cells can be broadly classified into 2 categories – primitive and centred unit cells.
- What is a unit cell? (1)
  - Name the three types of centred unit cells. (1 %)
  - The unit cell dimension of a particular crystal system is  $a = b = c$  and  $\alpha = \beta = \gamma = 90^\circ$ . Identify the crystal system. (1)
  - Give one example for the above crystal system. (%) [SAY 2013]
- 12) a) Every substance has some magnetic properties associated with it. How will you account for the following magnetic properties?
- Paramagnetic property (1)
  - Ferromagnetic property (1)
- b) A compound is formed by two elements P and Q. Atoms of Q (as anions) make hcp lattice and those of the element P (as cations) occupy all the tetrahedral voids. What is the formula of the compound? (2) [March 2014]
- 13) a) Crystalline solids are 'anisotropic'. What is anisotropy? (1)
- b) Copper crystals have fcc unit cells.
- Compute the number of atoms per unit cell of copper crystals. (1)
  - Calculate the mass of a unit cell of copper crystals. (Atomic mass of Cu = 63.54u) (2) [SAY 2014]
- 14) Unit cells can be divided into two categories – primitive and centred unit cells.
- Differentiate between unit cell and crystal lattice. (2)
  - Calculate the number of atoms per unit cell in the following:
    - Body centred cubic unit cell (bcc)
    - Face centred cubic unit cell (fcc) (2) [March 2015]
- 15) a) Which of the following is not a characteristic of a crystalline solid?
- Definite heat of fusion
  - isotropic nature
  - A regular orderly arrangement of constituent particles
  - A true solid (1)
- b) Frenkel defect and Schottky defects are two stoichiometric defects found in crystalline solids.
- What are stoichiometric defects? (1)
  - Write any two differences between Schottky and Frenkel defects? (2) [SAY 2015]
- 16) a) Which of the following is a molecular solid?
- (a) Diamond (b) graphite (c) ice (d) quartz (1)
- (b) Unit cells can be classified into primitive and centered unit cells. Differentiate between primitive and centered unit cells. (1)
- (c) Presence of excess Na makes NaCl crystal coloured. Explain on the basis of crystal defects. (2) [March 2016]
- 17) A unit cell is a term related to crystal structure.
- What do you mean by unit cell? (1)
  - Name any two types of cubic unit cells. (1)
  - Calculate the number of atoms in each of the above mentioned cubic unit cells. (1)
  - Identify the substance which shows Frenkel defect:
    - NaCl
    - KCl
    - ZnS
    - AgBr (1) [SAY 2016]
- 18) a) Identify the non-stoichiometric defect
- Schottky defect
  - Frenkel defect
  - interstitial defect
  - Metal deficiency defect (1)
- b) What type of substance could make better permanent magnets – ferrimagnetic or ferromagnetic? Justify your answer. (2)

# Guide For Plus Two Chemistry

**Terry C. Jones**



## **Guide For Plus Two Chemistry:**

## Reviewing **Guide For Plus Two Chemistry**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Guide For Plus Two Chemistry**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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