

2. Metallic magnesium reacts with steam to produce magnesium hydroxide and hydrogen gas.

a. Write a balanced equation for this reaction. (Note: It is very important that you do this step correctly!)



b. What type of reaction is this? single replacement

c. If 16.2 g of Mg are heated with 12.0 g of H_2O , which is the limiting reactant?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 24.0 \text{ g H}_2\text{O}$$

There is less H_2O (12.0 g) in our initial amount so
 $\boxed{\text{H}_2\text{O is the L.R}}$

d. How many moles of the excess reactant remain after the reaction is complete?

$$16.2 \text{ g Mg} \times \frac{1 \text{ mol Mg}}{24.309 \text{ g Mg}} \times \frac{2 \text{ mol H}_2\text{O}}{1 \text{ mol Mg}} = 1.3 \text{ mol H}_2\text{O}$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.666 \text{ mol H}_2\text{O}$$

$$1.3 \text{ mol H}_2\text{O} - 0.666 \text{ mol H}_2\text{O} = 0.634 \text{ mol H}_2\text{O} \times \frac{1 \text{ mol Mg}}{2 \text{ mol H}_2\text{O}} = \boxed{0.317 \text{ mol Mg}}$$

e. How many grams of each product are formed?

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol Mg}(\text{OH})_2}{2 \text{ mol H}_2\text{O}} \times \frac{58.32 \text{ g Mg}(\text{OH})_2}{1 \text{ mol Mg}(\text{OH})_2} = 19.4 \text{ g Mg}(\text{OH})_2$$

$$12.0 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} \times \frac{2.02 \text{ g H}_2}{1 \text{ mol H}_2} = 6.28 \text{ g H}_2$$

e. You perform this reaction, and obtain 16.0 g $\text{Mg}(\text{OH})_2$. What is your percent yield?

$$\% \text{ yield} = \frac{\text{actual} \times 100}{\text{theoretical}}$$

$$\frac{16.0 \text{ g Mg}(\text{OH})_2}{19.4 \text{ g Mg}(\text{OH})_2} \times 100 = 82.5\%$$

Gas Stoichiometry Problems And Answers

Jessica J Manson



Gas Stoichiometry Problems And Answers:

Dr. Salm's Chemistry Problem Drill Book Robert J. Salm, Robert J. Salm Robert J. Salm Ph.D., 2005 This drill book contains many common problem types that are asked in General Chemistry classes in High School and College This work will give you practice with the major problem types as you prepare for finals and standardized tests **Investigating Chemistry**

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