

Behavior of Gases Test Review

1. Fill the following table of the basic gas laws.

Gas Law	Equation	Variable held constant	Acronym
Boyle's Law	$P_1V_1 = P_2V_2$	Temperature	BaT
Charles's Law	$\frac{V_1}{T_1} = \frac{V_2}{T_2}$	Pressure	CaP
Gay Lussac's Law	$\frac{P_1}{T_1} = \frac{P_2}{T_2}$	Volume	GLoVe

2. What does ideal gas law allow you to calculate that the other gas laws do not? **moles**
3. What is the device that measures atmospheric pressure? **barometer**
4. Describe the following effects of a gas
- If temperature is held constant, what happens to the volume if pressure is increased?
if pressure is increased, the volume decreases
 - At a fixed volume, what happens to the particles if temperature is decreased? (i.e. do they move faster or slower?) What happens to pressure?
Particles move slower as the temperature is decreased and the pressure decreases
 - At a fixed pressure, what happens to the volume of a gas if temperature is increased?
The volume increases as the temperature increases
5. What is the standard temperature and pressure? (STP)
Standard temperature = 273 K Standard pressure = 1 atmosphere
6. How many liters does one mole of a gas occupy at STP?
1 mole = 22.4 Liters @ STP
7. Convert 0.56 atm to mmHg
 $0.56 \text{ atm} \times \frac{760 \text{ mmHg}}{1 \text{ atm}} = 425.6 \text{ mmHg}$
8. Convert 790 mmHg to kPa
 $790 \text{ mmHg} \times \frac{1 \text{ atm}}{760 \text{ mmHg}} \times \frac{101.3 \text{ kPa}}{1 \text{ atm}} = 105 \text{ kPa}$
9. Convert -22.6 °C to Kelvin
 $K = ^\circ\text{C} + 273$ $K = -22.6 + 273 = 250.4 \text{ K}$

Gas Laws Review Sheet Answers

Victor M. Corman

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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