

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

- 1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres (1.00 atm = 101.3 kPa) and a temperature of 340 K?
- 2) If I have a 50.0 liter container that holds 45.0 moles of gas at a temperature of 200° C, what is the pressure inside the container?
- 3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2.0 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?
- 4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 20.0 atmospheres?
- 5) I have a balloon that can hold 10.0 liters of air. If I blow up this balloon with 0.470 moles of helium gas at a pressure of 120 kPa, at what temperature (°C) is the balloon?
- 6) What is the molar mass of a gas if 10.0 grams of it occupy 2.87 dm³ at a temperature of 25.0°C and 196 kPa?
- 7) What is the density (in g/L) of Argon gas when the temperature is 35.°C and the pressure is 1.50 atm?

Gas Laws Practice Problems With Solutions

Arshad Iqbal



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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics. The focus is not to recap or review the theoretical concepts well described in the available texts. Instead, the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry. Explanations, instructional process sequences, solved examples, and completely solved practice problems are greatly expanded, containing significantly more detail than can usually be devoted to in a comprehensive text. Many chapters also provide alternative viewpoints as an aid to understanding. Key Features: The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester. Each is written in a specific and detailed step by step process for problem solving, whether mathematical or conceptual. Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts. Includes a chapter designed to eliminate confusion concerning acid base reactions, which often persists through working with acid base equilibrium. Many chapters provide alternative viewpoints as an aid to understanding. This book addresses a very real need for a large number of incoming freshmen in STEM fields.

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