

Gas Laws Problem Packet

Gas Laws Packet

Name _____ Date _____ Period _____

Instructions

Standard Conditions

Unit	Standard Pressure
atm	1.013
torr	760
mm	760
cm	760
mm Hg	760

22.4 L = 1 mol at STP

STP (Standard Temperature and Pressure) = 0°C and 1 atm

The gas laws:

Boyle's Law	Charles's Law
$P_1V_1 = P_2V_2$	$V_1/T_1 = V_2/T_2$
Combined Gas Law	Avogadro's Law
$P_1V_1/T_1 = P_2V_2/T_2$	$V_1/n_1 = V_2/n_2$
Avogadro's Law	Combined Gas Law
$V_1/n_1 = V_2/n_2$	$P_1V_1/T_1 = P_2V_2/T_2$
Combined Gas Law	Avogadro's Law
$P_1V_1/T_1 = P_2V_2/T_2$	$V_1/n_1 = V_2/n_2$
Avogadro's Law	Combined Gas Law
$V_1/n_1 = V_2/n_2$	$P_1V_1/T_1 = P_2V_2/T_2$

The constant R (used in the ideal gas law) depends on the unit of pressure:

- If atm, $R = 0.0821 \frac{\text{L} \cdot \text{atm}}{\text{mol} \cdot \text{K}}$
- If mm, $R = 62.36 \frac{\text{L} \cdot \text{mm Hg}}{\text{mol} \cdot \text{K}}$

Notes:

- Temperature must always be in Kelvin! $K = ^\circ C + 273$
- If you are given two different units of pressure, CONVERT (one).
- Create two columns before solving: initial and final. Write what you're given in the problem. Find the law that uses the variables you're given, and solve for the unknown. (for Ideal Gas Law, write what you are given, insert those in equation and solve for unknown)
- Avogadro's Law: DON'T FORGET TO USE MOLE! (Mol/L or mol/L)
- Graham's Law: You will ALWAYS have the molecular mass for these problems. Double and triple check you have the mole of effusion and molecular mass in the correct spots before solving the equation (this is an INVERSE relationship). You MUST calculate the square root of double both sides of the equation before moving things.

Charles' Law $V_1/T_1 = V_2/T_2$

1. The temperature under my refrigerator is about 40°F. If I place a balloon in the refrigerator, will it expand or contract? Why?

2. A balloon is at 20°C and 1 atm. If I move it to 0°C and 1 atm, what happens to its volume?

Boyle's Law $P_1V_1 = P_2V_2$

1. A gas is at 1 atm and 1 L. If I increase the pressure to 2 atm, what happens to the volume?

2. A gas is at 1 atm and 1 L. If I decrease the pressure to 0.5 atm, what happens to the volume?

Combined Gas Law $P_1V_1/T_1 = P_2V_2/T_2$

1. A gas is at 1 atm, 1 L, and 273 K. If I change the pressure to 2 atm and the temperature to 546 K, what happens to the volume?

2. A gas is at 1 atm, 1 L, and 273 K. If I change the pressure to 0.5 atm and the temperature to 546 K, what happens to the volume?

Avogadro's Law $V_1/n_1 = V_2/n_2$

1. A gas is at 1 atm, 1 L, and 273 K. If I change the number of moles to 2 moles, what happens to the volume?

2. A gas is at 1 atm, 1 L, and 273 K. If I change the number of moles to 0.5 moles, what happens to the volume?

Graham's Law $r_1/r_2 = \sqrt{M_2/M_1}$

1. Gas A has a molecular mass of 4 g/mol. Gas B has a molecular mass of 16 g/mol. What is the ratio of their effusion rates?

2. Gas A has a molecular mass of 4 g/mol. Gas B has a molecular mass of 16 g/mol. What is the ratio of their diffusion rates?

All Gas Laws Included!

Gas Law Problems Packet

SA Adler



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Unveiling the Power of Verbal Beauty: An Mental Sojourn through **Gas Law Problems Packet**

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