

1. A gas has a pressure of 0.370 atm at 55.0 °C. What is the pressure at standard temperature?

$$P_2 = P_1 \frac{T_2}{T_1} = \frac{(0.370 \text{ atm})(273 \text{ K})}{328 \text{ K}} = 0.313 \text{ atm}$$

2. If 2.00 mole of gas occupies 4.50L at STP. How much of the same gas will occupy 3.00L at STP?

$$n_2 = n_1 \frac{V_2}{V_1} = \frac{(2.00 \text{ mole})(3.00 \text{ L})}{(4.50 \text{ L})} = 1.33 \text{ mole}$$

3. I've got a car with an internal volume of 12,000 L. If I drive my car into the river and it implodes, what will be the volume of the gas when the pressure goes from 1.0 atm to 1.4 atm?

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{(1.0 \text{ atm})(12,000 \text{ L})}{(1.4 \text{ atm})} = 8571 \text{ L}$$

4. What's the partial pressure of carbon dioxide in a container that holds 5.00 moles of carbon dioxide, 3.00 moles of nitrogen, and 1.00 mole of hydrogen and has a total pressure of 1.05 atm?

$$P_{\text{CO}_2} = P_T \times \frac{\text{mole CO}_2}{\text{Total moles}} = (1.05 \text{ atm}) \times \frac{(5.00 \text{ moles})}{(9.00 \text{ moles})} = 0.583 \text{ atm}$$

5. I have 130 liters of gas in a piston at a temperature of 250. °C. If I cool the gas until the volume decreases to 85 liters, what will temperature of the gas be?

$$T_2 = \frac{V_2 T_1}{V_1} = \frac{(85 \text{ L})(250 \text{ K})}{130 \text{ L}} = 162 \text{ K} = -111 \text{ °C}$$

6. If a gas in a closed container is pressurized from 15.0 atmospheres to 16.0 atmospheres and its original temperature was 25.0 °C, what would the final temperature of the gas be?

$$T_2 = \frac{P_2 T_1}{P_1} = \frac{(16.0 \text{ atm})(298 \text{ K})}{(15.0 \text{ atm})} = 318 \text{ K}$$

7. To compress nitrogen at 1 atm from 750 mL to 500 mL, what must the new pressure be if the temperature is kept constant?

$$P_2 = \frac{V_1 P_1}{V_2} = \frac{(750 \text{ mL})(1 \text{ atm})}{(500 \text{ mL})} = 1.5 \text{ atm}$$

8. Calcium carbonate decomposes at 1200. °C to form carbon dioxide and calcium oxide. If 25 liters of carbon dioxide are collected at 1200. °C, what will the volume of this gas be after it cools to 25 °C?

$$V_2 = \frac{V_1 T_2}{T_1} = \frac{(25 \text{ L})(298 \text{ K})}{(1473 \text{ K})} = 5.1 \text{ L}$$

9. 5.00 L of a gas is known to contain 0.065 moles. If this amount of gas is increased to 1.80 moles, what new volume will result (at an unchanged temperature and pressure)?

$$V_2 = \frac{V_1 n_2}{n_1} = \frac{(5.00 \text{ L})(1.80 \text{ moles})}{(0.065 \text{ moles})} = 138 \text{ L}$$

Gas Laws Worksheet 1 Answer Key

Clifford Lane

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Table of Contents Gas Laws Worksheet 1 Answer Key

1. Understanding the eBook Gas Laws Worksheet 1 Answer Key
 - The Rise of Digital Reading Gas Laws Worksheet 1 Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Worksheet 1 Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Laws Worksheet 1 Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Worksheet 1 Answer Key
 - Personalized Recommendations
 - Gas Laws Worksheet 1 Answer Key User Reviews and Ratings
 - Gas Laws Worksheet 1 Answer Key and Bestseller Lists
5. Accessing Gas Laws Worksheet 1 Answer Key Free and Paid eBooks
 - Gas Laws Worksheet 1 Answer Key Public Domain eBooks

- Gas Laws Worksheet 1 Answer Key eBook Subscription Services
- Gas Laws Worksheet 1 Answer Key Budget-Friendly Options
- 6. Navigating Gas Laws Worksheet 1 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Laws Worksheet 1 Answer Key Compatibility with Devices
 - Gas Laws Worksheet 1 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Laws Worksheet 1 Answer Key
 - Highlighting and Note-Taking Gas Laws Worksheet 1 Answer Key
 - Interactive Elements Gas Laws Worksheet 1 Answer Key
- 8. Staying Engaged with Gas Laws Worksheet 1 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Laws Worksheet 1 Answer Key
- 9. Balancing eBooks and Physical Books Gas Laws Worksheet 1 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Laws Worksheet 1 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Laws Worksheet 1 Answer Key
 - Setting Reading Goals Gas Laws Worksheet 1 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Laws Worksheet 1 Answer Key
 - Fact-Checking eBook Content of Gas Laws Worksheet 1 Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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