

# DO NOT WRITE

## Gas Laws Problems Worksheet(alternate assignment to note activity) Show all work including equations and units to receive full credit !!

**Charles' Law: What remains constant?** \_\_\_\_\_

1. A gas occupies 3.7 L at 100.0 C. What volume will it occupy at 50.0 C?

**Boyle's Law: What remains constant?** \_\_\_\_\_

2. A gas occupies 5.0 L at 101.3 kPa. What volume will it occupy at 1.7 atm?

3. A gas occupies 10.0 L at 1.2 atm. What pressure in kPa will cause it to expand to 15.0 L?

**Combined gas law: What remains constant?** \_\_\_\_\_

4. At what temperature (in C) will a sample of gas occupy 25.0 L with a pressure of 215 kPa if the sample occupies 50.0 L at 13.5 C and 105 kPa?

5. A gas sample occupies 50.0L at 13.5 C and 105 kPa. What volume should it occupy at 750 K and 75.3 kPa?

6. At what pressure (in atm) will a gas sample occupy 30.0 L with a temperature of 273 K if it occupies 50.0 L at 13.5 C and 105 kPa.?

**Ideal gas law:**

7. How many liters would 2.7 moles of an ideal gas occupy if it were at 20.0 C and 1.3 atm ?

8. Suppose two 200.0 L tanks are to be filled separately with the gases helium and hydrogen. What mass of each gas is needed to produce a pressure of 135 atm in its respective tank at 24 C?

9. What volume is occupied by 5.03 g of O<sub>2</sub> at 28 C and a pressure of 0.725 atm?

**Dalton's law of partial pressures:**

10. (a) Write out Dalton's law of partial pressures here in equation form:

(b) State Avogadro's hypothesis here:

11. A gas mixture of Helium and Nitrogen is at 273 K and 200 kPa. If the pressure in the tank due to the presence of the Nitrogen gas is 123 kPa, what is the pressure from the Helium?  
What percent of the gas particles in the tank are Nitrogen molecules?

12. If 2.43 % of the particles in a sample of air are Argon, what is the partial pressure of Ar in a 1.00 L sample of air at STP (101.3 kPa and 0.0 C)?

13. Mixtures of helium and oxygen are used in scuba diving tanks to help prevent 'the bends.' For a particular dive, 46 L of O<sub>2</sub> at 25 C and 1.00 atm was pumped along with 12 L of He at 25 C and 1.00 atm into a tank with a volume of 5.0 L . Calculate the partial pressure of each gas *and* the total pressure in the tank at 25 C.

# DO NOT WRITE

# Gas Laws And Scuba Diving Worksheet Answer Key

**Hussin A.Rothana**



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## **Table of Contents Gas Laws And Scuba Diving Worksheet Answer Key**

1. Understanding the eBook Gas Laws And Scuba Diving Worksheet Answer Key
  - The Rise of Digital Reading Gas Laws And Scuba Diving Worksheet Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws And Scuba Diving Worksheet Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Laws And Scuba Diving Worksheet Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws And Scuba Diving Worksheet Answer Key
  - Personalized Recommendations
  - Gas Laws And Scuba Diving Worksheet Answer Key User Reviews and Ratings
  - Gas Laws And Scuba Diving Worksheet Answer Key and Bestseller Lists

5. Accessing Gas Laws And Scuba Diving Worksheet Answer Key Free and Paid eBooks
  - Gas Laws And Scuba Diving Worksheet Answer Key Public Domain eBooks
  - Gas Laws And Scuba Diving Worksheet Answer Key eBook Subscription Services
  - Gas Laws And Scuba Diving Worksheet Answer Key Budget-Friendly Options
6. Navigating Gas Laws And Scuba Diving Worksheet Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Laws And Scuba Diving Worksheet Answer Key Compatibility with Devices
  - Gas Laws And Scuba Diving Worksheet Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Laws And Scuba Diving Worksheet Answer Key
  - Highlighting and Note-Taking Gas Laws And Scuba Diving Worksheet Answer Key
  - Interactive Elements Gas Laws And Scuba Diving Worksheet Answer Key
8. Staying Engaged with Gas Laws And Scuba Diving Worksheet Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Laws And Scuba Diving Worksheet Answer Key
9. Balancing eBooks and Physical Books Gas Laws And Scuba Diving Worksheet Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Laws And Scuba Diving Worksheet Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws And Scuba Diving Worksheet Answer Key
  - Setting Reading Goals Gas Laws And Scuba Diving Worksheet Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws And Scuba Diving Worksheet Answer Key
  - Fact-Checking eBook Content of Gas Laws And Scuba Diving Worksheet Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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