

Name \_\_\_\_\_

## Gas Laws

21. Q: What is the relationship between volume and moles in Avogadro's Principle?

A: Volume is directly proportional to the number of moles.

22. Q: Calculate the pressure of 1.0 mole of gas in a 10.0 L container at 273 K using the ideal gas law.

A: The pressure is 2.24 atm.

23. Q: According to the ideal gas law, what happens to the pressure if the volume is reduced by half while the temperature and number of moles remain constant?

A: The pressure doubles.

24. Q: If the volume of a gas is tripled and the temperature is doubled, what happens to the pressure according to the combined gas law?

A: The pressure decreases to 2/3 of its original value.

25. Q: Using gas stoichiometry, how many liters of  $\text{CO}_2$  are produced from 5.0 L of  $\text{O}_2$  in the reaction  $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$  at STP?

A: 5.0 L of  $\text{CO}_2$  are produced.

26. Q: In the reaction  $2\text{NO} + \text{O}_2 \rightarrow 2\text{NO}_2$ , how many liters of  $\text{NO}_2$  are produced from 4.0 L of  $\text{NO}$  at STP?

A: 4.0 L of  $\text{NO}_2$  are produced.

27. Q: Calculate the volume of  $\text{N}_2$  gas required to produce 20.0 L of  $\text{NH}_3$  in the reaction  $\text{N}_2 + 3\text{H}_2 \rightarrow 2\text{NH}_3$  at STP.

A: 10.0 L of  $\text{N}_2$  are required.

28. Q: If the temperature of a gas is increased from 300 K to 600 K and the volume is kept constant, what happens to the pressure?

A: The pressure doubles.

29. Q: Using the ideal gas law, how do you calculate the volume of gas if you know the pressure, temperature, and number of moles?

A:  $V = nRT/P$

30. Q: In the reaction  $\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$ , how many liters of  $\text{HCl}$  are produced from 3.0 L of  $\text{H}_2$  at STP?

A: 6.0 L of  $\text{HCl}$  are produced.



# Gas Law Study Guide Answers

**J Elliott**



## **Gas Law Study Guide Answers:**

## Adopting the Tune of Term: An Emotional Symphony within **Gas Law Study Guide Answers**

In a world eaten by displays and the ceaseless chatter of instant transmission, the melodic elegance and psychological symphony produced by the prepared term usually diminish into the backdrop, eclipsed by the constant noise and disruptions that permeate our lives. However, set within the pages of **Gas Law Study Guide Answers** an enchanting literary value overflowing with organic feelings, lies an immersive symphony waiting to be embraced. Crafted by an elegant musician of language, this fascinating masterpiece conducts viewers on a psychological trip, skillfully unraveling the hidden melodies and profound impact resonating within each cautiously crafted phrase. Within the depths of the emotional analysis, we can explore the book's key harmonies, analyze its enthralling writing model, and surrender ourselves to the profound resonance that echoes in the depths of readers' souls.

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