

Chapter 14—WS GAS LAWS

1. A quantity of gas under a pressure of 800 mm Hg has a volume of 380 liters. What is its volume at standard pressure when temperature is held constant?
2. A quantity of gas has a volume of 120 liters when confined under a pressure of 700.0 mm Hg at a temperature of 20.0°C. At what pressure will its volume be 30.0 liters at 20.0°C?
3. At constant pressure, the volume of a gas is increased from 150 liters to 300.0 liters by heating it. If the original temperature of the gas was 20.0°C, what will its final temperature be?
4. A quantity of gas exerts a pressure of 740 mm Hg at a temperature of 22.0°C. If the volume remains unchanged, what pressure will it exert at -8.0°C?
5. A quantity of oxygen gas has a volume of 850 ml when measured at 27.0°C and 730 mm Hg of pressure. Determine its volume at STP.
6. When measured at STP, a quantity of gas has a volume of 500.0 liters. What volume will it occupy at 0.0°C and a pressure of 700.0 mm Hg?
7. Know the relationships between your four variables- pressure, volume, temperature, and moles. Which are inverse and which are direct?? Be able to explain them in terms of what is happening to the energy, collisions, and space between the particles.
8. A gas whose behavior closely resembles that of an ideal gas has a volume of 2.00 liters at a temperature of 27.0°C and a pressure of 900.0 mm Hg. (a) How many moles of molecules are in the sample? (b) How many molecules are in the sample?
9. Know Graham's law of effusion. Be able to calculate the ratio of rate of effusion. Example: what is the rate of effusion of oxygen gas compared to nitrogen if both gases are at the same temperature?

10. How many moles?

$$1. P_1 V_1 = P_2 V_2$$

$$800 \text{ mmHg} (380 \text{ L}) = 760 \text{ mmHg} (V_2)$$

$$V_2 = 400 \text{ L}$$

$P \downarrow V \uparrow$ checks out!

$$2. P_1 V_1 = P_2 V_2$$

$$700.0 \text{ mmHg} (120 \text{ L}) = P_2 (30 \text{ L})$$

$$P_2 = 2800 \text{ mmHg}$$

$V \downarrow P \uparrow$ checks out!

$$3. \frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \frac{150 \text{ L}}{293 \text{ K}} = \frac{300 \text{ L}}{T_2}$$

$$T_2 (150) = 300 (293)$$

$$T_2 = 586 \text{ K}$$

$V \uparrow T \uparrow$
checks out!

Gas Law Review Sheet Answers

**Barron's Educational Series, Brian
Stewart**



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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Table of Contents Gas Law Review Sheet Answers

1. Understanding the eBook Gas Law Review Sheet Answers
 - The Rise of Digital Reading Gas Law Review Sheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Review Sheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Law Review Sheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Review Sheet Answers
 - Personalized Recommendations

- Gas Law Review Sheet Answers User Reviews and Ratings
- Gas Law Review Sheet Answers and Bestseller Lists
- 5. Accessing Gas Law Review Sheet Answers Free and Paid eBooks
 - Gas Law Review Sheet Answers Public Domain eBooks
 - Gas Law Review Sheet Answers eBook Subscription Services
 - Gas Law Review Sheet Answers Budget-Friendly Options
- 6. Navigating Gas Law Review Sheet Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Law Review Sheet Answers Compatibility with Devices
 - Gas Law Review Sheet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Law Review Sheet Answers
 - Highlighting and Note-Taking Gas Law Review Sheet Answers
 - Interactive Elements Gas Law Review Sheet Answers
- 8. Staying Engaged with Gas Law Review Sheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Law Review Sheet Answers
- 9. Balancing eBooks and Physical Books Gas Law Review Sheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Law Review Sheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gas Law Review Sheet Answers
 - Setting Reading Goals Gas Law Review Sheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gas Law Review Sheet Answers
 - Fact-Checking eBook Content of Gas Law Review Sheet Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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