

Gas Stoichiometry Exam  
2016-2017

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Multiple Choice

Refer to the set of graphs below when answering the next two questions:

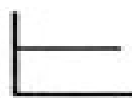
Graph #1



Graph #2



Graph #3



Graph #4



- B 1. Which of the graphs above is the best qualitative graph of pressure versus volume of nitrogen at constant temperature and number of moles?

A) Graph #1  
B) Graph #2  
C) Graph #3  
D) Graph #4

→ INVERSE

- A 2. Which of the graphs above is the best qualitative graph of Kelvin temperature versus pressure of an ideal gas at constant volume and number of moles?

A) Graph #1  
B) Graph #2  
C) Graph #3  
D) Graph #4

→ DIRECT

- D 3. Use the kinetic molecular theory of gases to predict what would happen to a closed sample of a gas whose temperature increased while its volume decreased.

→ MOVE FASTER, MORE COLLISIONS  
A) The average kinetic energy of the molecules of the gas would decrease.  
B) Its pressure would decrease.  
C) The number of moles of gas would decrease.  
D) Its pressure would increase.

- B 4. A hydrocarbon gas with the empirical formula  $\text{CH}_2$  has a density of 1.3 g/L at  $0^\circ\text{C}$  and 1.00 atm. A possible formula for the hydrocarbon is:

A)  $\text{CH}_2$   
B)  $\text{C}_2\text{H}_4$   
C)  $\text{C}_3\text{H}_6$   
D)  $\text{C}_4\text{H}_8$   
E)  $\text{C}_5\text{H}_{10}$

~~1.3 g/L~~

$$1.3 \text{ g/L} \times 22.4 \text{ L/mol} \approx 30 \text{ g/mol}$$

# Gases Test B Answer Key

**Rex M. Heyworth**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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