

CHEM 101 Worksheet, Chapter 5: Gases

Multiple choice questions 1-10: circle the correct answer(A-E)

Show your work questions 11-14: fill in the answer

1 atm = 760 mm Hg = 101,325 Pa 0 °C = 273 K R = 0.08206 L.atm/(mol.K) = 8.314 J/(mol.K)

1. 2.67×10^5 Pa = _____ mm Hg?

- A. 3.56×10^6 mm Hg B. 3.47×10^{-7} mm Hg C. 3.51 mm Hg D. 20.0 mm Hg E. 200 mm Hg

2. How many pascal (Pa) is 7.5×10^{-2} mm Hg?

- A. 100 Pa B. 5.78×10^4 Pa C. 10.0 Pa D. 9.97×10^{-7} Pa E. 1000 Pa

3. What is the volume occupied by 3.36 g $\text{NH}_3(\text{g})$ when the pressure is 740 mm Hg and its temperature 27 °C?

- A. 5.00 L B. 4.74 L C. 1.44×10^3 L D. 0.117 L E. 500 L

4. What is the mass of CO_2 in a 340. mL container when its pressure is 2.80 atm and the temperature is 57 °C?

- A. 26.7 g B. 0.125 g C. 125 g D. 3.52 g E. 3.47×10^{-4} g

5. A certain amount of gas at 300K is held in a volume of 1.0L. When the gas is heated to 600K (the volume is still 1.0 L) the pressure will:

- A. stay the same B. be four times as high C. be double the initial pressure
D. be one half the initial pressure E. be one fourth of the initial pressure

6. Which of the following gases will have the highest density, when all are at the same temperature and pressure?

- A. $\text{CO}_2(\text{g})$ B. $\text{NO}_2(\text{g})$ C. $\text{C}_3\text{H}_8(\text{g})$ D. $\text{SO}_2(\text{g})$ E. $\text{N}_2\text{O}_4(\text{g})$

7. Calculate the density of butane, $\text{C}_4\text{H}_{10}(\text{g})$ at p = 800 mm Hg and T = 37 °C.

- A. 2.17 g/L B. 0.272 g/L C. 2.17 g/m³ D. E. 0.0236 g/L E. 2.40 g/L

8. An unknown gas has a density of 2.67 g/L at p = 1.0 atm and T = 19 °C. What is the molar mass of the gas?

- A. 64.0 g/mol B. 5.76 g/L C. 640 g/mol D. 64. 0 kg/mol E. 32.0 g.mol

9. What volume of $\text{CO}_2(\text{g})$ will be formed when 2.0 g $\text{CaCO}_3(\text{s})$ is heated and the gas is collected at 700 mm pressure and 20 °C? The reaction is: $\text{CaCO}_3(\text{s}) \rightarrow \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$

- A. 0.44 L B. 0.52 L C. 52 L D. 3.61 L E. 22.4 L

10. A mixture of 3.0 mol $\text{H}_2(\text{g})$ and 2.0 mol $\text{N}_2(\text{g})$ has a total pressure of 4.5 atm. What are the partial pressures of H_2 and N_2 in the mixture?

- A. 2.5 atm H_2 , 2.0 atm N_2 B. 0.5 atm H_2 , 4.0 atm N_2 C. 1.8 atm H_2 , 2.7 atm N_2
D. 4.0 atm H_2 , 0.5 atm N_2 E. 2.7 atm H_2 , 1.8 atm N_2

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