13.	Consider the following statements							
	i) Atmospheric pressure is less at the top of a mountain than at sea level							
	ii) Gases are much more compressible than solids or liquids							
	iii) When the atmospheric pressure increases the height of the mercury column rises Select the correct statement							
	a) I and II	and II b) II at		nd III c) I		ш	d) I, II	and III
14.	Compressibility factor for $\rm CO_2$ at 400 K and 71.0 bar is 0.8697. The molar volume of $\rm CO_2$ under these conditions is							
	a) 22.04 dm ³ b) 2.24) 2.24 dm ³	dm³ c) 0.41 dm³		d) 19.5	d) 19.5dm³	
15.	If temperature and volume of an ideal gas is increased to twice its values, the initial pressure P becomes							
	a) 4P	Ь) 2P		c) P		d) 3P	
16.	At identical temperature and pressure, the rate of diffusion of hydrogen gas is $3\sqrt{3}$ times that of a hydrocarbon having molecular formula C_nH_{2n-2} . What is the value of n?							
100	8	b) 4	c) 3		d) I			
	Equal moles of hydrogen and oxygen gases are placed in a container, with a pin-hole through which both can escape what fraction of oxygen escapes in the time required for one-half of the hydrogen to escape. (NEET phase I)							
a	3/8	b) ½	c)]	8	d) 1/4			
1.8.	The variation of volume V, with temperature T, keeping pressure constant is called the							
	coefficient of thermal expansion ie $\alpha = \frac{1}{V} \left(\frac{\partial V}{\partial T} \right)_{p}$. For an ideal gas α is equal to							
(a)	T	b) ½	c) P	ta i	d) none	of these		
19.		ants) are i	n the order	r Q < R <	S < P. At			(a, b are Vander ature, among the
(a)) P	b) Q	c) R		d) S			
20.	Maximum de	eviation fr	rom ideal g	as is expe	cted from	n (NEET)		
a) 21	CH ₄ (g) The units of	b) NH, Vander W	(g) aals consta		g) d'a' respe		(g)	
24.)	mol L-1 and L mol L and L	b) mol L and L atm mol ² d) none of these						

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Chandresh Agrawal, Nandini Books

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