

6 • The Gas Laws

G A S S T O I C H I O M E T R Y W O R K S H E E T

Please answer the following on separate paper using proper units and showing all work. Please note that these problems require a **balanced chemical equation**.

- Carbon monoxide reacts with oxygen to produce carbon dioxide. If 1.0 L of carbon monoxide reacts with oxygen at STP,
 - How many liters of oxygen are required to react?
 - How many liters of carbon dioxide are produced?
- Acetylene gas (C_2H_2) undergoes combustion to produce carbon dioxide and water vapor,
 - How many liters of C_2H_2 are required to produce 75.0 L of CO_2 ?
 - What volume of H_2O is produced?
 - What volume of O_2 is required?
- If liquid carbon disulfide (CS_2) reacts with 450 mL of oxygen to produce the gases carbon dioxide and sulfur dioxide, what volume of each product is produced?
- Assume that 5.60 L of hydrogen gas at STP reacts with copper (II) oxide according to the following balanced equation:
$$\text{CuO (s)} + \text{H}_2 \text{ (g)} \rightarrow \text{Cu (s)} + \text{H}_2\text{O (g)}$$
 - How many moles of H_2 react?
 - How many moles of copper are produced?
 - How many grams of copper are produced?

Gas Stoichiometry Problem And Answer

John T. Moore



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