

Gas Law Stoichiometry Worksheet

Directions: Use significant figures and units in the problems below.

$$PV = nRT$$

1. Given the following unbalanced chemical equation for the combination reaction of sodium metal and chlorine gas:



- a. What volume of chlorine gas, measured at STP, is necessary for the complete reaction of 4.81 g of sodium metal.

$$4.81 \text{ g Na} \left| \frac{1 \text{ mol}}{23.0 \text{ g}} \right| \left| \frac{1 \text{ mol Cl}_2}{2 \text{ mol Na}} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mol}} \right| = 2.34 \text{ L Cl}_2$$

$$2.34 \text{ L Cl}_2$$



- a. Balance the above reaction.
b. What volume of oxygen gas at 25°C and 1.04 atm is needed for the complete combustion of 5.53 g of propane?

$$n = x$$

$$V = 0.628 \text{ L}$$

$$T_1 = 25 + 273 = 298 \text{ K}$$

$$P_1 = 1.05 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.628 \text{ mol}) (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (298 \text{ K})}{1.05 \text{ atm}} = 14.6 \text{ L}$$

3. Potassium permanganate, KMnO_4 , is produced commercially by oxidizing aqueous potassium manganate, K_2MnO_4 .



- a. Balance the above reaction.
b. What volume of $\text{Cl}_2(g)$, measured at STP, is needed to produce 10.0 g of KMnO_4 ?

$$V = \frac{nRT}{P}$$

$$10.0 \text{ g KMnO}_4 \left| \frac{1 \text{ mole}}{158 \text{ g}} \right| \left| \frac{1 \text{ mole Cl}_2}{2 \text{ moles KMnO}_4} \right| \left| \frac{22.4 \text{ L}}{1 \text{ mole Cl}_2} \right| = 0.709 \text{ L Cl}_2$$

4. If water is added to magnesium nitride, ammonia gas is produced when the mixture is heated.



- a. Balance the above reaction.
b. If 10.3 g of magnesium nitride is treated with 10.3 g of water, what volume of ammonia gas would be collected at 24°C and 752 mmHg?

$$10.3 \text{ g Mg}_3\text{N}_2 \left| \frac{1 \text{ mole}}{100.9 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{1 \text{ mole}} \right| = 0.204 \text{ mol}$$

$$10.3 \text{ g H}_2\text{O} \left| \frac{1 \text{ mole}}{18 \text{ g}} \right| \left| \frac{2 \text{ NH}_3}{3 \text{ moles H}_2\text{O}} \right| = 0.381 \text{ mol}$$

$$PV = nRT$$

$$24 + 273 = 297 \text{ K}$$

$$752 / 760 = 0.989 \text{ atm}$$

$$V = \frac{nRT}{P} = \frac{(0.204 \text{ mol}) (0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}) (297 \text{ K})}{0.989 \text{ atm}} = 5.03 \text{ L}$$

Gas Stoichiometry Answer Key

RJ Shavelson

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Gas Stoichiometry Answer Key:

Decoding **Gas Stoichiometry Answer Key**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Gas Stoichiometry Answer Key**," a mesmerizing literary creation penned by way of a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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