

12 • The Gas Laws**CHARLES'S LAW**

Charles' Law states the volume of a gas varies directly with the Kelvin temperature, assuming the pressure is constant. We use the following formulas:

$$\frac{V_1}{T_1} = \frac{V_2}{T_2} \quad \text{or} \quad V_1 \times T_2 = V_2 \times T_1$$

$$K = ^\circ C + 273$$

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Solve the following problems assuming a constant pressure. Assume all numbers are 3 significant figures.

1. A sample of nitrogen occupies a volume of 250 mL at $25^\circ C$. What volume will it occupy at $95^\circ C$?
 $V_1 = 250 \text{ mL}$, $T_1 = 298 \text{ K}$, $T_2 = 368 \text{ K}$
2. Oxygen gas is at a temperature of $40^\circ C$ when it occupies a volume of 2.30 Liters. To what temperature should it be raised to occupy a volume of 6.50 Liters?
 $T_1 = 313 \text{ K}$, $V_1 = 2.30 \text{ L}$, $V_2 = 6.50 \text{ L}$
3. Hydrogen gas was cooled from $150^\circ C$ to $50^\circ C$. Its new volume is 75.0 mL. What was its original volume?
 $T_1 = 423 \text{ K}$, $T_2 = 323 \text{ K}$, $V_2 = 75.0 \text{ mL}$
4. Chlorine gas occupies a volume of 25.0 mL at 300 K. What volume will it occupy at 600 K?
 $V_1 = 25.0 \text{ mL}$, $T_1 = 300 \text{ K}$, $T_2 = 600 \text{ K}$
5. A sample of neon gas at $50^\circ C$ and a volume of 2.50 Liters is cooled to $25^\circ C$. What is the new volume?
 $T_1 = 323 \text{ K}$, $V_1 = 2.50 \text{ L}$, $T_2 = 298 \text{ K}$
6. Fluorine gas at 300 K occupies a volume of 500 mL. To what temperature should it be lowered to bring the volume to 300 mL?
 $T_1 = 300 \text{ K}$, $V_1 = 500 \text{ mL}$, $V_2 = 300 \text{ mL}$
7. Helium occupies a volume of 3.80 Liters at $-45^\circ C$. What volume will it occupy at $45^\circ C$?
 $V_1 = 3.80 \text{ L}$, $T_1 = 228 \text{ K}$, $T_2 = 318 \text{ K}$
8. A sample of argon gas is cooled and its volume went from 380 mL to 250 mL. If its final temperature was $-55^\circ C$, what was its original temperature?
 $V_1 = 380 \text{ mL}$, $V_2 = 250 \text{ mL}$, $T_2 = 218 \text{ K}$

Gas Laws Worksheet Boyle Charles And Combined Gas Laws Answers

Lydia Anderson



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