

**Gas Laws Worksheet**

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

**Boyle's Law Problems:**  $P_1 V_1 = P_2 V_2$ 

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mmHg})(22.5 \text{ L}) = (725 \text{ mmHg}) V_2$$

$$V_2 = \frac{(748 \text{ mmHg})(22.5 \text{ L})}{(725 \text{ mmHg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

# Gas Laws Practice Problems With Answers

**A Loxley**



## **Gas Laws Practice Problems With Answers:**

This book delves into Gas Laws Practice Problems With Answers. Gas Laws Practice Problems With Answers is a vital topic that needs to be grasped by everyone, ranging from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Gas Laws Practice Problems With Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
    - Chapter 1: Introduction to Gas Laws Practice Problems With Answers
    - Chapter 2: Essential Elements of Gas Laws Practice Problems With Answers
    - Chapter 3: Gas Laws Practice Problems With Answers in Everyday Life
    - Chapter 4: Gas Laws Practice Problems With Answers in Specific Contexts
    - Chapter 5: Conclusion
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- This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Gas Laws Practice Problems With Answers.

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## **Table of Contents Gas Laws Practice Problems With Answers**

1. Understanding the eBook Gas Laws Practice Problems With Answers
  - The Rise of Digital Reading Gas Laws Practice Problems With Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Gas Laws Practice Problems With Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Gas Laws Practice Problems With Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Laws Practice Problems With Answers
  - Personalized Recommendations
  - Gas Laws Practice Problems With Answers User Reviews and Ratings
  - Gas Laws Practice Problems With Answers and Bestseller Lists
5. Accessing Gas Laws Practice Problems With Answers Free and Paid eBooks
  - Gas Laws Practice Problems With Answers Public Domain eBooks
  - Gas Laws Practice Problems With Answers eBook Subscription Services
  - Gas Laws Practice Problems With Answers Budget-Friendly Options
6. Navigating Gas Laws Practice Problems With Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Gas Laws Practice Problems With Answers Compatibility with Devices
  - Gas Laws Practice Problems With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Gas Laws Practice Problems With Answers
  - Highlighting and Note-Taking Gas Laws Practice Problems With Answers
  - Interactive Elements Gas Laws Practice Problems With Answers

8. Staying Engaged with Gas Laws Practice Problems With Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Gas Laws Practice Problems With Answers
9. Balancing eBooks and Physical Books Gas Laws Practice Problems With Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Gas Laws Practice Problems With Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Gas Laws Practice Problems With Answers
  - Setting Reading Goals Gas Laws Practice Problems With Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Laws Practice Problems With Answers
  - Fact-Checking eBook Content of Gas Laws Practice Problems With Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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

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