

Gas Laws Worksheet

$$\text{atm} = 760.0 \text{ mm Hg} = 101.3 \text{ kPa} = 760.0 \text{ torr}$$

Boyle's Law Problems:

$$P_1 V_1 = P_2 V_2$$

1. If 22.5 L of nitrogen at 748 mm Hg are compressed to 725 mm Hg at constant temperature. What is the new volume?

$$(748 \text{ mm Hg})(22.5 \text{ L}) = (725 \text{ mm Hg}) V_2$$

$$V_2 = \frac{(748 \text{ mm Hg})(22.5 \text{ L})}{(725 \text{ mm Hg})}$$

$$V_2 = 23.2 \text{ L}$$

2. A gas with a volume of 4.0 L at a pressure of 205 kPa is allowed to expand to a volume of 12.0 L. What is the pressure in the container if the temperature remains constant?

$$(4.0 \text{ L})(205 \text{ kPa}) = (12.0 \text{ L}) P_2$$

$$P_2 = \frac{(4.0 \text{ L})(205 \text{ kPa})}{12.0 \text{ L}}$$

$$P_2 = 68.3 \text{ kPa}$$

3. What pressure is required to compress 196.0 liters of air at 1.00 atmosphere into a cylinder whose volume is 26.0 liters?

$$(196.0 \text{ L})(1.00 \text{ atm}) = (26.0 \text{ L}) P_2$$

$$P_2 = \frac{(196.0 \text{ L})(1.00 \text{ atm})}{26.0 \text{ L}}$$

$$P_2 = 7.54 \text{ atm}$$

4. A 40.0 L tank of ammonia has a pressure of 12.7 kPa. Calculate the volume of the ammonia if its pressure is changed to 8.4 kPa while its temperature remains constant.

$$(40.0 \text{ L})(12.7 \text{ kPa}) = (8.4 \text{ kPa}) V_2$$

$$V_2 = \frac{(40.0 \text{ L})(12.7 \text{ kPa})}{8.4 \text{ kPa}}$$

$$V_2 = 60.5 \text{ L}$$

Gas Law Practice Homework Packet

Jianjun Gao



Gas Law Practice Homework Packet:

Reviewing **Gas Law Practice Homework Packet**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is actually astonishing. Within the pages of "**Gas Law Practice Homework Packet**," an enthralling opus penned by a very acclaimed wordsmith, readers set about an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve into the book's central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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Table of Contents Gas Law Practice Homework Packet

1. Understanding the eBook Gas Law Practice Homework Packet
 - The Rise of Digital Reading Gas Law Practice Homework Packet
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Law Practice Homework Packet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Law Practice Homework Packet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Law Practice Homework Packet
 - Personalized Recommendations
 - Gas Law Practice Homework Packet User Reviews and Ratings
 - Gas Law Practice Homework Packet and Bestseller Lists

5. Accessing Gas Law Practice Homework Packet Free and Paid eBooks
 - Gas Law Practice Homework Packet Public Domain eBooks
 - Gas Law Practice Homework Packet eBook Subscription Services
 - Gas Law Practice Homework Packet Budget-Friendly Options
6. Navigating Gas Law Practice Homework Packet eBook Formats
 - ePub, PDF, MOBI, and More
 - Gas Law Practice Homework Packet Compatibility with Devices
 - Gas Law Practice Homework Packet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gas Law Practice Homework Packet
 - Highlighting and Note-Taking Gas Law Practice Homework Packet
 - Interactive Elements Gas Law Practice Homework Packet
8. Staying Engaged with Gas Law Practice Homework Packet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Law Practice Homework Packet
9. Balancing eBooks and Physical Books Gas Law Practice Homework Packet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Law Practice Homework Packet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Law Practice Homework Packet
 - Setting Reading Goals Gas Law Practice Homework Packet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Law Practice Homework Packet
 - Fact-Checking eBook Content of Gas Law Practice Homework Packet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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