

Gas Stoichiometry Practice 1.0 - Answer Key**Objective:** - Perform stoichiometry calculations of gases including those at STP using molar volume**Directions:** Solve each gas stoichiometry problem. Show all work. Include units.

1. Calculate the volume of hydrogen gas that will react with exactly 4.6 L of oxygen gas using the equation below. Assume there is no change in temperature or pressure.



$$4.6 \text{ L O}_2 \left(\frac{2 \text{ L H}_2}{1 \text{ L O}_2} \right) = \boxed{9.2 \text{ L H}_2}$$

2. What volume of CO_2 gas will be produced from 17.5 L of oxygen gas reacting with excess ethane?



$$17.5 \text{ L O}_2 \left(\frac{4 \text{ L CO}_2}{7 \text{ L O}_2} \right) = \boxed{10.0 \text{ L CO}_2}$$

3. What volume of hydrogen at STP can be produced from the reaction of 36.5 g of zinc with hydrochloric acid?



$$36.5 \text{ g Zn} \left(\frac{1 \text{ mol Zn}}{65.38 \text{ g Zn}} \right) \left(\frac{1 \text{ mol H}_2}{1 \text{ mol Zn}} \right) \left(\frac{22.4 \text{ L H}_2}{1 \text{ mol H}_2} \right) = \boxed{12.5 \text{ L H}_2}$$

4. How many grams of NaCl can be produced by the reaction of 12.6 L of chlorine at STP with an excess of sodium?



$$12.6 \text{ L Cl}_2 \left(\frac{1 \text{ mol Cl}_2}{22.4 \text{ L Cl}_2} \right) \left(\frac{2 \text{ mol NaCl}}{1 \text{ mol Cl}_2} \right) \left(\frac{58.44 \text{ g NaCl}}{1 \text{ mol NaCl}} \right) = \boxed{65.7 \text{ g NaCl}}$$

5. In order to produce 50.0 liters of ammonia gas (NH_3) at STP, how many grams of both hydrogen and nitrogen would you need to react?



$$50.0 \text{ L NH}_3 \left(\frac{1 \text{ mol NH}_3}{22.4 \text{ L NH}_3} \right) \left(\frac{1 \text{ mol N}_2}{2 \text{ mol NH}_3} \right) \left(\frac{28.02 \text{ g N}_2}{1 \text{ mol N}_2} \right) = \boxed{21.3 \text{ g N}_2}$$

$$50.0 \text{ L NH}_3 \left(\frac{1 \text{ mol NH}_3}{22.4 \text{ L NH}_3} \right) \left(\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} \right) \left(\frac{2.02 \text{ g H}_2}{1 \text{ mol H}_2} \right) = \boxed{6.75 \text{ g H}_2}$$

Gas Stoichiometry Practice With Answers

Matthew Johll



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Table of Contents Gas Stoichiometry Practice With Answers

1. Understanding the eBook Gas Stoichiometry Practice With Answers
 - The Rise of Digital Reading Gas Stoichiometry Practice With Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Gas Stoichiometry Practice With Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gas Stoichiometry Practice With Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gas Stoichiometry Practice With Answers
 - Personalized Recommendations
 - Gas Stoichiometry Practice With Answers User Reviews and Ratings
 - Gas Stoichiometry Practice With Answers and Bestseller Lists
5. Accessing Gas Stoichiometry Practice With Answers Free and Paid eBooks

- Gas Stoichiometry Practice With Answers Public Domain eBooks
 - Gas Stoichiometry Practice With Answers eBook Subscription Services
 - Gas Stoichiometry Practice With Answers Budget-Friendly Options
6. Navigating Gas Stoichiometry Practice With Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Gas Stoichiometry Practice With Answers Compatibility with Devices
 - Gas Stoichiometry Practice With Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Gas Stoichiometry Practice With Answers
 - Highlighting and Note-Taking Gas Stoichiometry Practice With Answers
 - Interactive Elements Gas Stoichiometry Practice With Answers
8. Staying Engaged with Gas Stoichiometry Practice With Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gas Stoichiometry Practice With Answers
9. Balancing eBooks and Physical Books Gas Stoichiometry Practice With Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gas Stoichiometry Practice With Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gas Stoichiometry Practice With Answers
- Setting Reading Goals Gas Stoichiometry Practice With Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gas Stoichiometry Practice With Answers
- Fact-Checking eBook Content of Gas Stoichiometry Practice With Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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






















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