

Gas Laws Study Guide

Notes:

Properties of solids, liquids, & gases:

Solids	Liquids	Gases
○ Particles very close together ○ Highest density ○ Lowest kinetic energy ○ Least amount of motion ○ Definite shape & volume	○ Particles loosely held together ○ Moderate density ○ Moderate kinetic energy ○ Some motion ○ No definite shape, definite volume	○ Particles very far apart ○ Least dense ○ Most kinetic energy ○ High amount of motion ○ No definite shape or volume ○ No organization

Phase Changes

Solid → liquid: melting

Liquid → solid: freezing

gas → liquid: condensation

liquid → gas: evaporating/vaporizing

gas → solid: deposition

solid → gas: sublimation

Gases:

There are three main characteristics of gases – all of which are a result of the kinetic molecular theory which says:

1. All particles of a gas are in constant motion
2. As a result of this motion, random collisions occur, but gases are not attracted or repelled from each other
3. Gases are much smaller than the distance between them.

The characteristics of gases are:

1. Volume: the amount of space a gas occupies – usually measured in liters (L) or milliliters (mL)
2. Pressure: force per unit area – comes from particles banging into the side of the container. Measured in kilopascals, atmospheres, or mm of mercury.
3. Temperature: average kinetic energy of the particles of gas – measured in degrees Celsius or Kelvin.
4. STP: standard temperature & pressure

****ALWAYS HAS TO BE IN KELVIN FOR GAS LAW PROBLEMS****

Pressure: 1 atm = 101.3 kPa = 760 mmHg = 760 torr

Temperature: 0°C = 273 K K = °C + 273

Volume: 1 mole = 22.4 L

Scientific Laws & Discoveries:

1. **Dalton's Law of Partial Pressure:** the total pressure of a gas is equal to the sum of the pressures of each part of the gas mixture.

$$P_{\text{total}} = P_1 + P_2 + P_3 + \dots + P_n$$
2. **Graham's Law of Effusion:** at constant temperature and pressure, the rate of effusion is inversely proportional to the square root of their molar masses.

$$\text{rate of effusion}_{\text{gas 1}} \propto \sqrt{\text{molar mass}_{\text{gas 1}}} \quad \text{rate of effusion}_{\text{gas 2}} \propto \sqrt{\text{molar mass}_{\text{gas 2}}}$$
3. **Boyle's Law:** if the temperature remains the same, pressure & volume are indirectly proportional. If one condition goes up, the other goes down & vice versa.

$$P_1 V_1 = P_2 V_2$$
4. **Charles's Law:** if the pressure remains constant, the temperature & volume are directly proportional. If one condition goes up, so does the other.

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$
5. **Gay-Lussac's Law:** if the volume remains constant, the pressure & temperature are directly proportional – if one condition goes up, so does the other.

$$\frac{P_1}{T_1} = \frac{P_2}{T_2}$$
6. **Combined Gas Law:** shows the previous 3 gas laws in one. To use any of the 3 previous laws, simply eliminate the variable that remains constant.

$$\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$$
7. **Ideal Gas Law:** Avogadro came along and said that perhaps the number of particles (atoms or molecules) in the gas sample may affect the three characteristics of gases and he was right. So, the number of moles (n) in the sample also has an effect. You may have to convert from grams to moles, moles to molecules, or vice versa depending on what the question gives/wants.

$PV = nRT$ R is the gas law constant (there are 3 different ones you can use based on the unit the pressure is given in).

R = 0.0821 L•atm
mol•K

8.31 L•kPa
mol•K

62.4 L•mmHg
mol•K

****Be careful to make sure the units cancel!!!**

Gas Laws Answer Study Guide Chemistry

Konstantinos Papadopoulos



Gas Laws Answer Study Guide Chemistry:

Gas Laws Answer Study Guide Chemistry Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become much more apparent than ever. Its capability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Gas Laws Answer Study Guide Chemistry**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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web aug 19 2022 bunun yerine netflix uyarlama için sürpriz bir yeni bölüm duyurdu bir basın açıklamasıyla duyuran netflix iki bölüm bonus bölümün neil gaiman ın ikonik karanlık fantastik roman serisi the sandman dream country nin üçüncü cildinden hikayeler içereceğini doğruladı

the sandman tv dizisini internetten izleyin justwatch - Feb 09 2023

web gözlerimizi kapatıp uykuya daldığımızda bizi bambaşka bir dünya bekler düşdünya denen bu yerde düşler lordu sandman en derindeki tüm korkularımızı ve fantezilerimizi şekillendirir

netflix in bugüne kadarki en maliyetli dizisi the sandman - Jun 01 2022

web 05 08 2022 17 45 dc dünyası yeni bir diziye imza atıyor Çizgi romanlardan tanıdığımız sandman artık bu çizgi romanlardan çıkıp ekranlarımıza geliyor netflix in tam tamına 165 milyon dolar bütçe ayırdığı the sandman dizisi hakkında bilmeniz gerekenleri bu içeriğimizde yer verdik

the sandmand dizisi konusu nedir netflix dizisi the sandman - Jul 02 2022

web aug 23 2022 8 ağustos 2022 tarihinde izleyiciyle buluşan netflix in en yüksek bütçeli olan dizisi olarak duyurulan the sandman dizisinin konusu ve oyuncularına ilişkin bilgileri sizler için bu

the sandman dizi 2022 sinemalar com - Apr 30 2022

web Çağdaş kurgu tarihi dramının kusursuz bir şekilde iç içe geçtiği modern mit ve karanlık fantazinin zengin bir karışımı olan the sandman rüya kralı morpheus tan etkilenen insanları ve yerleri takip ederken onun yaptığı kozmik ve insani hataları anlatıyor

the sandman rotten tomatoes - Sep 04 2022

web aug 5 2022 when the sandman aka dream the cosmic being who controls all dreams is captured and held prisoner for

more than a century he must journey across different worlds and timelines to fix the chaos