

# Properties of Acids and Bases

## Identification and Classification

### Introduction

Acids and bases are useful reagents in the chemistry laboratory, and they play an important role in biology and nature. What are acids and bases? What properties can be used to distinguish between them? Let's explore the properties of acids and bases and identify the characteristic features that will allow us to classify substances as acids and bases.

### Concepts

- Acids and bases
- Indicators
- pH scale
- Neutralization
- Active metals
- Conductivity

### Background

The word *acid* is derived from the Latin verb *acere*, which means “to (be) sour.” The origin of the word itself reveals a characteristic physical property of acids—they taste sour. Lemons, oranges and grapefruits are called *citrus* fruits because they contain citric acid, an acidic compound that gives them their sour taste. Although taste is an interesting property of the food we eat, it is not a property that we will use in the chemistry laboratory to classify compounds as acids or bases. The following properties are typically used to classify compounds as acids or bases.

- **Conductivity.** Some acids and bases ionize completely when dissolved in water. Solutions that contain large numbers of dissolved ions conduct an electric current and are called *electrolytes*. Other weaker acids and bases may ionize only partially when dissolved in water and may conduct electricity only weakly—they are called *weak electrolytes*. Substances that do not produce dissolved ions will not conduct electricity and are called *nonelectrolytes*.
- **Reactions with Metals.** Acids react with so-called *active metals*—reactive metals, such as magnesium and zinc—to produce hydrogen gas and solutions of metal ions. The reactions of different metals with acids is a well-known test used to rank metals from most active to least active. Reaction of a single active metal with a variety of different solutions is one of the best methods to identify acids and compare their relative acidity.

- **Effect on Indicators.** Indicators are organic dyes that change color in acidic or basic solutions. One of the oldest known acid–base indicators is litmus, a natural dye obtained from lichens. Its use was described as early as the 16th century. Litmus paper, prepared by soaking paper in a solution of the dye, is often used as a general test for acids and bases. Phenolphthalein is another indicator that shows a color change as solutions change from acidic to basic. Although these indicators are useful for broadly classifying substances as acids or bases, they are not able to distinguish among different levels of acidity or basicity. By using combinations of different indicators, however, it is possible to obtain a spectrum of color changes over a wide range of acidity levels. Universal indicator and pH paper are two products that use combinations of indicators to rank substances from most acidic to least acidic or most basic to least basic.

- **The pH Scale.** The pH scale is a numerical scale used to describe the relative acidity or basicity of a solution related to the concentration of  $\text{H}_3\text{O}^+$  ions. The abbreviation pH stands for “power of hydrogen”—a difference of one unit on the pH scale corresponds to a power of ten difference in the concentration of  $\text{H}_3\text{O}^+$  ions. Pure water contains extremely small but equal concentrations of  $\text{H}_3\text{O}^+$  and  $\text{OH}^-$  ions due to self-ionization (Equation 1). Acids and bases are substances that alter the concentrations of  $\text{H}_3\text{O}^+$  and  $\text{OH}^-$  ions in solution.



- **Neutralization Reactions.** Acids and bases react with each other to give *neutral* products—solutions that are neither acidic nor basic. The products of the neutralization of an acid and a base are an ionic compound (generally referred to as a salt) and water. The amount of acid that will react with a specific amount of base is governed by stoichiometry. Neutralization reactions are frequently used in the laboratory to determine how much of an acidic or basic compound is present in a substance.

### Experiment Overview

The purpose of this experiment is to explore the properties of aqueous solutions and to classify them as acidic, basic or neutral. The results will be used to develop working definitions and to study the pH scale for identifying acids and bases.

# Flinn Chemtopic Labs Properties Of Acids And Bases Answers

**H. L. Finston, Allen C. Rychtman**



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**Acid Base Equilibria** E. King, R. Robinson, 1965-01-01 **It's an Acid Test! Acids vs Bases | Understanding the Properties of Acids and Bases | Grade 6-8 Physical Science** Baby Professor, 2024-04-15 Dive into the intriguing world of acids and bases with this essential guide for middle school students Perfect for teachers homeschooling parents and librarians this book demystifies the sour taste of lemons and the bitterness of soap through an engaging exploration of acidity and basicity With chapters dedicated to understanding the properties the pH scale and the everyday applications of acids and bases it is a vital addition to any STEM curriculum sparking curiosity and foundational knowledge in young scientists *New Solid Acids and Bases* K. Tanabe, M. Misono, H. Hattori, Y. Ono, 1990-02-02 This volume summarises and reviews the enormous progress made over the past two decades in solid acids and bases with emphasis on fundamental aspects and chemical principles In recent years many new kinds of solid acids and bases have been found and synthesized The surface properties in particular acidic and basic properties and the structures of the new solids have been clarified by newly developed measurement methods using modern instruments and techniques The characterized solid acids and bases have been applied as catalysts for diversified reactions many good correlations being obtained between the acid base properties and the catalytic activities or selectivities Recently acid base bifunctional catalysis on solid surfaces is becoming a more and more important and intriguing field of study It has been recognized that the acidic and basic properties of catalysts and catalyst supports play an important role in oxidation reduction hydrogenation hydrocracking etc The effect of the preparation method

and the pretreatment conditions of solid acids and bases on the acidic and basic properties the nature of acidic and basic sites and the mechanism regarding the generation of acidity and basicity have been elucidated experimentally and theoretically On the basis of the accumulated knowledge of solid acids and bases it is now possible to design and develop highly active and selective solid acid and base catalysts for particular reactions The chemistry of solid acids and bases is now being related to and utilized in numerous areas including adsorbents sensors cosmetics fuel cells sensitized pressed papers and others The information presented in this book will therefore be of interest to a wide ranging readership Solid Acids and Bases Kozo Tanabe,2012-12-02 Solid Acids and Bases Their Catalytic Properties reviews developments in the studies of acidic and basic properties of solids including the efficacy and special characteristics of solid acid and base catalysts This book discusses the determination of basic and acidic properties on solid surfaces and relationship between acid strength and acid amount The structure and acid base properties of mixed metal oxides and correlation between acid base properties and catalytic activity and selectivity are also deliberated This publication is useful to professional chemists and graduate students in the fields of organic inorganic and physical chemistry petroleum chemistry and catalysis including readers interested in the acidic and basic properties on solid surfaces Acid-base Indicators Izaak Maurits Kolthoff,1937 The dissociation of strong and weak electrolytes The properties of acid base indicators The colorimetric determination of hydrogen ion concentration **A New View of Current Acid-base Theories** H. L. Finston,Allen C. Rychtmann,1982 **Acids and Bases** Chris Oxlade,2001-08-29 What is the difference between a base and an alkali How do acids react with metals What does the pH scale measure Many people think that chemistry is something that is done by scientists in laboratories full of test tubes and flasks of colored liquids But chemistry does not just happen in the lab It happens all around us at home at school and everywhere else The books in the Chemicals in Action series will help you discover how chemistry works and how it is used in our everyday lives Explore the microscopic world of atoms and molecules find out about the chemicals in our bodies and perform experiments to learn more Read Acids Bases to find out what gives acids and bases their properties how they react with each other and how we use them in our everyday lives You will also find several experiments that can be done at home Book jacket New Solid Acids and Bases ,1989 **Strong Acids and Bases-General Properties** ,2016 You return to the topic of pH and learn about how pH relates to two kinds of compounds acids and bases Through an introductory problem you explore the relationship of various ions within these compounds **New Solid Acids and Bases** Kōzō Tanabe,1989 **Acid-base Equilibria** Edward Jasper King,1965 **Acids & Bases** Chris Oxlade,2002 This book discusses what gives acids and bases their properties how acids and bases react with each other and how we use them in our everyday lives **Acids, Bases and Salts** Brian Knapp,1998-01-01 Provides photographic and textual demonstrations of a range of chemistry tests which explore the properties of acids bases and salts Suggested level secondary

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