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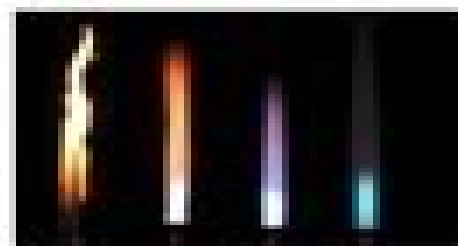
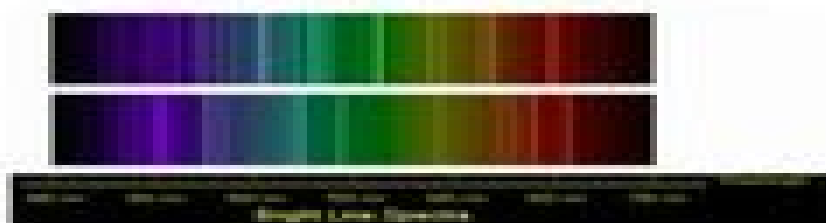
## Flame Test Lab

**Essential Questions:** What happens when different cations are exposed to a flame?

Reading Questions	Class	Class	Evidence	Reasoning	Model
/10	/10	/10	/5	/5	/10

**Materials:**

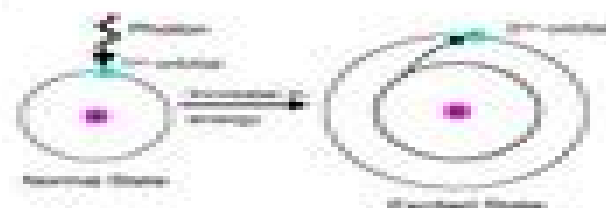
- 1) Bunsen Burner    2) Q-tips    3) Flint lighter    4) Various c Chlorides    5) 100 mL beaker



### Background Information:

If you excite an atom or an ion by very strong heating, electrons can be promoted from their normal unexcited state into higher orbitals. As they fall back down to lower levels (either in one go or in several steps), energy is released as light. Each of these jumps involves a specific amount of energy being released as light energy and each corresponds to a particular wavelength (or frequency). As a result of all these jumps, a spectrum of lines will be produced, some of which will be in the visible part of the spectrum. The color you see will be a combination of all these individual colors.

In the case of sodium (or other metalloids), the jumps involve very high energies and these result in lines in the UV part of the spectrum which your eyes can't see. The jumps that you can see in flame tests come from electrons falling from a higher to a lower level in the metal atoms. So if, for example, you put sodium chloride which contains sodium ions, into a flame, where do the atoms come from? In the hot flame, some of the sodium ions regain their electrons to form neutral sodium atoms again. A sodium atom in an unexcited state has the structure 1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>1</sup>, but within the flame there will be all sorts of excited states of the electrons. Sodium's familiar bright orange-yellow flame color results from promoted electrons falling back from the 3p<sup>1</sup> level to their normal 3s<sup>1</sup> level.<sup>1</sup>



<sup>1</sup>University of York, 2019, May 15. Flame Tests. Retrieved from <https://www.youtube.com/watch?v=388888888888>

# Flame Test And Electron Configuration Lab Answer

**Julie Gess-Newsome, Julie Luft, Randy L.  
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**ChemDiscovery Teacher Edition** Olga I. Agapova, 2002      **Chemistry** Eugene LeMay, Jr., Herbert Beall, Karen M. Robblee, Douglas C. Brower, Prentice-Hall Staff, 2002-02      Scientific and Technical Aerospace Reports , 1992      **Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998      **Nuclear Science Abstracts** , 1967      *Popular Science* , 1979-10 Popular Science gives our readers the information and tools to improve their technology and their world The core belief that Popular Science and our readers share The future is going to be better and science and technology are the driving forces that will help make it better      *Energy: a Continuing Bibliography with Indexes* , 1977      U.S. Government Research Reports , 1962      **Applied Mechanics Reviews** , 1995      Energy Research Abstracts , 1989 Semiannual with semiannual and annual indexes References to all scientific and technical literature coming from DOE its laboratories energy centers and contractors Includes all works deriving from DOE other related government sponsored information and foreign nonnuclear information Arranged under 39 categories e g Biomedical sciences basic studies Biomedical sciences applied studies Health and safety and Fusion energy Entry gives bibliographical information and abstract Corporate author subject report number indexes      **Holt Chemistry** Ralph Thomas Myers, 2004      *ERDA Energy Research Abstracts* United States. Energy Research and Development Administration, 1976-05      ERDA Energy Research Abstracts United States. Energy Research and Development Administration. Technical Information Center, 1976      *An Introductory Guide to EC Competition Law and Practice* Valentine Korah, 1994      **Technical Information Pilot** , 1949

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