

Qualitative analysis of an organic compound showed that it contained only carbon, hydrogen and oxygen. A quantitative study of the same compound was performed, in which a 0.5096 g sample was burnt in excess oxygen to produce 0.4160 g of water and 700.7 mL of carbon dioxide, collected at 100.0 °C and 102.8 kPa.

373K

(a) Determine the empirical formula of the compound.

(10 marks)

C	H	O
0.7007 L of CO_2 $n(\text{C}) = n(\text{CO}_2)$ $PV = nRT \rightarrow n = \frac{PV}{RT}$ $n(\text{CO}_2) = \frac{102.8 \times 0.7007}{8.315 \times 373}$ $= 2.3218 \times 10^{-2} \text{ mol}$ $n(\text{C}) = nM = 0.27885 \text{ g}$	0.416 g of H_2O $n = \frac{m}{M} = \frac{0.416}{18.016} = 2.3091 \times 10^{-2} \text{ mol}$	$m(\text{O}) = 0.5096 - m(\text{C}) - m(\text{H})$ $= 0.18423 \text{ g}$ $n = \frac{m}{M} = \frac{0.18423}{16} = 1.1514 \times 10^{-2}$

EMPIRICAL FORMULA

EXAM QUESTION

$$\frac{2.3218}{1.1514} = 2$$

Empirical Formula Gravimetric Questions Answers

Clifford Lane



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