

# Experiment 7 Report Sheet

## Empirical Formulas

Date \_\_\_\_\_ Lab Sec \_\_\_\_\_ Name \_\_\_\_\_

Indicate whether the following data are for Part B, Part C, or Part D. \_\_\_\_\_

Desk No. \_\_\_\_\_

Identify the original sample \_\_\_\_\_

	Trial 1	Trial 2
1. Mass of crucible and lid (g)	25.636	40.347
2. Mass of crucible, lid, and sample (g)	25.786	40.542
3. Mass of sample (g) 0.150	0.150	0.195
4. Instructor's approval of flame and apparatus	—	—
5. Mass of crucible, lid, and product		
1st mass measurement (g)	25.858	40.654
2nd mass measurement (g)	25.856	40.652
3rd mass measurement (g)	25.859	40.658
6. Final mass of crucible, lid, and product (g)	25.860	40.655
7. Mass of product (g)	0.224g	0.113g
8. Part B. Combination Reaction of Magnesium and Oxygen		
a. Mass of O (g)	0.014g	
b. Mass ratio of Mg to O	2:1	
c. Moles of Mg (mol)	0.00617 mol	
d. Moles of O (mol)	0.00463 mol	
e. Mole ratio of Mg to O	1:1	
f. Consensus empirical formula of magnesium oxide	1 MgO	
g. Percent by mass (%): 67.0 % Mg; 33.0 % O		

### 9. Part C. Decomposition Reaction of a Pure Compound

a. Mass of product = mass CaO (g)	0.224	
b. Mass of CO <sub>2</sub> (g)	0.1263	

e) note

# Empirical Formula Of A Compound Lab Report

**Albert A Gayle**



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